

Elementary Principles of Chemical Processes Solutions Manual 3rd Update Edition Edition Richard M. Felder

CHAPTER TWO

2.2 (a) $\frac{1.25 \text{ kg}}{1 \text{ min}} \left(\frac{1 \text{ hr}}{60 \text{ min}} \right) \left(\frac{1000 \text{ g}}{1 \text{ kg}} \right) = \frac{20.8 \text{ kg}}{1 \text{ hr}} \left(\frac{1000 \text{ g}}{1 \text{ kg}} \right) = 20,800 \text{ g/hr}$

(b) $\frac{0.5 \text{ kg} \cdot \text{hr}^{-1}}{1 \text{ hr}} \left(\frac{1000 \text{ g}}{1 \text{ kg}} \right) \left(\frac{1 \text{ min}}{60 \text{ sec}} \right) = \frac{8.33 \text{ g}}{1 \text{ min}} \left(\frac{1 \text{ sec}}{60 \text{ min}} \right) = 0.139 \text{ g/min}$

(c) $\frac{0.04 \text{ mol}}{1 \text{ kg}} \left(\frac{1 \text{ g}}{1000 \text{ kg}} \right) \left(\frac{1 \text{ hr}}{60 \text{ min}} \right) \left(\frac{1000 \text{ g}}{1 \text{ kg}} \right) = \frac{6.67 \times 10^{-5} \text{ mol}}{1 \text{ min}} \left(\frac{1 \text{ sec}}{60 \text{ min}} \right) = 1.11 \times 10^{-6} \text{ mol/sec}$

2.3 (a) $\frac{100 \text{ mol}}{1 \text{ hr}} \left(\frac{1 \text{ min}}{60 \text{ min}} \right) \left(\frac{1 \text{ kg}}{1000 \text{ mol}} \right) = \frac{1.67 \text{ kg}}{1 \text{ min}}$

(b) $\frac{100 \text{ kg}}{1 \text{ hr}} \left(\frac{1 \text{ min}}{60 \text{ min}} \right) \left(\frac{1 \text{ mol}}{100 \text{ kg}} \right) = \frac{0.167 \text{ mol}}{1 \text{ min}}$

(c) $\frac{1.27 \times 10^{-5} \text{ mol}}{1 \text{ kg}} \left(\frac{1 \text{ min}}{60 \text{ min}} \right) \left(\frac{1000 \text{ g}}{1 \text{ kg}} \right) \left(\frac{1 \text{ hr}}{60 \text{ min}} \right) = \frac{3.56 \times 10^{-7} \text{ mol}}{1 \text{ min}} \left(\frac{1 \text{ sec}}{60 \text{ min}} \right) = 5.93 \times 10^{-9} \text{ mol/sec}$

2.4 Assume that a golf ball (mass = 0.045 kg) is launched from a 2 m $\times$ 2 m $\times$ 2 m volume. If we assume a 1000 m/s launch velocity (2000 ft $\times$ 1000 ft $\times$ 1000 ft),

$m_{\text{ball}} = 0.045 \text{ kg} \times 10^3 \text{ m/s} = 45 \text{ kg} \cdot \text{m/s} = 100 \text{ lb} \cdot \text{ft/s}$

The estimate could vary by an order of magnitude or more, depending on the assumptions made.

2.5 (a) Right: $\frac{100 \text{ g}}{1 \text{ kg}} \left(\frac{1 \text{ min}}{60 \text{ min}} \right) \left(\frac{1000 \text{ g}}{1 \text{ kg}} \right) \left(\frac{1 \text{ hr}}{60 \text{ min}} \right) = \frac{2.78 \text{ g}}{1 \text{ min}} \left(\frac{1 \text{ sec}}{60 \text{ min}} \right) = 4.63 \times 10^{-3} \text{ g/sec}$

(b) Wrong: the units do not cancel to 1000000 units.

$\frac{1000000 \text{ g}}{1 \text{ kg}} \left(\frac{1 \text{ min}}{60 \text{ min}} \right) \left(\frac{1000 \text{ g}}{1 \text{ kg}} \right) \left(\frac{1 \text{ hr}}{60 \text{ min}} \right) = 2.78 \times 10^6 \text{ g/sec}$

2.6 (a) $\frac{1000 \text{ mol}}{1 \text{ kg}} \left(\frac{1 \text{ min}}{60 \text{ min}} \right) \left(\frac{1000 \text{ g}}{1 \text{ kg}} \right) \left(\frac{1 \text{ hr}}{60 \text{ min}} \right) = 2.78 \text{ g/sec}$

(b) $\frac{100 \text{ kg}}{1 \text{ hr}} \left(\frac{1 \text{ min}}{60 \text{ min}} \right) \left(\frac{1 \text{ mol}}{100 \text{ kg}} \right) = 0.167 \text{ mol/min}$

Calculate the total cost for 10000 units.

Total Cost (units) = $10000 \text{ units} \times \frac{0.167 \text{ mol}}{1 \text{ min}} \left(\frac{1 \text{ g}}{1000 \text{ mol}} \right) \left(\frac{1 \text{ min}}{60 \text{ min}} \right) \left(\frac{1 \text{ hr}}{60 \text{ min}} \right) = 10000 \text{ units} \times 2.78 \times 10^{-6} \text{ g/sec}$

Total Cost (units) = $10000 \text{ units} \times \frac{0.167 \text{ mol}}{1 \text{ min}} \left(\frac{1 \text{ g}}{1000 \text{ mol}} \right) \left(\frac{1 \text{ min}}{60 \text{ min}} \right) \left(\frac{1 \text{ hr}}{60 \text{ min}} \right) = 10000 \text{ units} \times 2.78 \times 10^{-6} \text{ g/sec}$

Repeat the two costs: $10000 \text{ units} \times 2.78 \times 10^{-6} \text{ g/sec}$

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