

# 8 • Electron Configurations & Periodicity

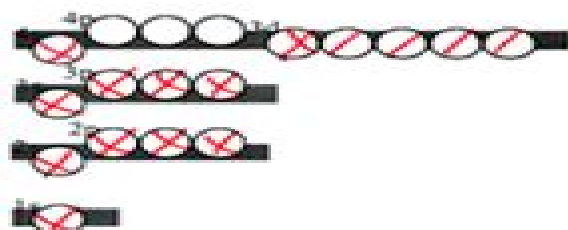
## WRITING ELECTRON CONFIGURATIONS

For each given element, fill in the orbital diagram and then write the electron configuration for the element.

|                                                 |                                                 |                                               |                                                |                                                |                                                 |
|-------------------------------------------------|-------------------------------------------------|-----------------------------------------------|------------------------------------------------|------------------------------------------------|-------------------------------------------------|
| 1.                                              | 2.                                              | 3.                                            | 4.                                             | 5.                                             | 6.                                              |
|                                                 |                                                 |                                               |                                                |                                                |                                                 |
| Element: Ar<br># of e <sup>-</sup> s: <u>18</u> | Element: Mg<br># of e <sup>-</sup> s: <u>12</u> | Element: N<br># of e <sup>-</sup> s: <u>7</u> | Element: Li<br># of e <sup>-</sup> s: <u>3</u> | Element: P<br># of e <sup>-</sup> s: <u>15</u> | Element: Cl<br># of e <sup>-</sup> s: <u>17</u> |

Write the electron configurations of each of these in **long form** and **short form**:

- Ar  $1s^2 2s^2 2p^6 3s^2 3p^4$   
 Ar  $[Ne] 3s^2 3p^4$
- Mg  $1s^2 2s^2 2p^6 3s^2$   
 Mg  $[Ne] 3s^2$
- N  $1s^2 2s^2 2p^3$   
 N  $[He] 2s^2 2p^3$
- Li  $1s^2 2s^1$   
 Li  $[He] 2s^1$
- P  $1s^2 2s^2 2p^4 3s^2 3p^3$   
 P  $[Ne] 3s^2 3p^3$
- Cl  $1s^2 2s^2 2p^6 3s^2 3p^5$   
 Cl  $[Ne] 3s^2 3p^5$



- Fill in the orbital diagram for the element, Fe, and write the electron configuration of Fe in the long and short form.



# Electron Configuration Practice Answer Key Table

**Richard Bailey**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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