

Electron Energy and Light

How does light reveal the behavior of electrons in an atom?

Why?

From fireworks to stars, the color of light is useful in finding out what's in matter. The emission of light by hydrogen and other atoms has played a key role in understanding the electronic structure of atoms. Trace materials, such as evidence from a crime scene, lead in paint or mercury in drinking water, can be identified by heating or burning the materials and examining the color(s) of light given off in the form of bright-line spectra.

Model 1 – White Light



1. Trace the arrows in Model 1 and shade in the table with colored pencils where appropriate.

2. What happens to white light when it passes through a prism?

The prism projects all colors of the rainbow.

3. Why are the color labels in the table in Model 1 plural (i.e., "Reds" rather than "Red")?

Because there are various types of each color depending on their wave length and photon energy.

4. Do all colors of light travel at the same speed?

yes

5. Do all colors of light have the same energy? If no, which colors have the highest energy and the least energy respectively?

no. Blues and violets have the most energy while oranges and reds have the least.



6. Consider the light illustrated in Model 1.

a. Which color corresponds to the longest wavelength?

Red

b. Which color corresponds to the shortest wavelength?

Violet

c. Write a sentence that describes the relationship between wavelength and energy of light.

The shorter the wavelength of light, the greater the energy of the light.

Electron Energy And Light Answer Sheet

Mark Steyvers

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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