

Stoichiometry "End of Chapter" IN CLASS Practice Assessment

Background Info -

- Why are the coefficients in a balanced equation necessary for stoichiometry?
Coefficients are necessary for establishing a molar/mole ratio to convert between moles of different substances.

Skill 1 - Using a balanced equation, calculate using mole to mole ratios. Use a balanced equation, given the grams of reactant, calc. the grams of product.

- Consider the following unbalanced equation:



There are 5.56g of manganese(IV) oxide present, calculate the grams of chlorine gas.

$$5.00\text{g MnO}_2 \times \frac{1\text{mol MnO}_2}{86.94\text{g MnO}_2} \times \frac{1\text{mol Cl}_2}{1\text{mol MnO}_2} \times \frac{70.90\text{g Cl}_2}{1\text{mol Cl}_2} = 4.00\text{g Cl}_2$$

Skill 2 - After calculating theoretical yield in grams, calculate the percent yield.

- You are assigned to complete a lab. Before the reaction is complete, you calculate that a reaction should produce 4.0g Cu(s). After the lab is over you find that 2.8g Cu(s) was collected. What would the percent yield be for that reaction completed within the lab?

$$\% \text{yield} = \left[\frac{\text{actual/theoretical}}{1} \right] \times 100 \\ = \left[\frac{2.8\text{g}/4.0\text{g}}{1} \right] \times 100 = 70\%$$

Skill 3 - Given experimental data, determine the limiting reactant and the excess reactant.

- Calcium carbonate combines with gaseous hydrogen chloride to produce solid calcium chloride, gaseous carbon dioxide, and water vapor. What is the percent yield if 155g of calcium carbonate combines with 250.0g of hydrogen chloride and only 142.0g of calcium chloride is actually obtained within the lab setting? Complete and show work for all 3 steps of a stoichiometry problem.



$$155\text{g CaCO}_3 \times \frac{1\text{mol CaCO}_3}{100.09\text{g CaCO}_3} = 1.55\text{mol CaCO}_3 \quad \text{L.R.} \\ 1.55\text{mol} \\ 250\text{g HCl} \times \frac{1\text{mol HCl}}{36.45\text{g HCl}} = 6.857\text{mol HCl} \quad \text{2mol in rxn} = 3.43\text{mol}$$

$$1.55\text{mol CaCO}_3 \times \left(\frac{1\text{mol CaCl}_2}{1\text{mol CaCO}_3} \right) \times \left(\frac{110.98\text{g CaCl}_2}{1\text{mol CaCl}_2} \right) = 172.02\text{g}$$

$$\% \text{yield} = \left[\frac{\text{actual/theoretical}}{1} \right] \times 100 \\ = \left[\frac{142.0\text{g}/172.02\text{g}}{1} \right] \times 100 = 82.55\%$$

Mixed Review Stoichiometry Answer Key

Rachel S Tattersall



Mixed Review Stoichiometry Answer Key:

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Table of Contents Mixed Review Stoichiometry Answer Key

1. Understanding the eBook Mixed Review Stoichiometry Answer Key
 - The Rise of Digital Reading Mixed Review Stoichiometry Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Mixed Review Stoichiometry Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in a Mixed Review Stoichiometry Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mixed Review Stoichiometry Answer Key
 - Personalized Recommendations
 - Mixed Review Stoichiometry Answer Key User Reviews and Ratings
 - Mixed Review Stoichiometry Answer Key and Bestseller Lists

5. Accessing Mixed Review Stoichiometry Answer Key Free and Paid eBooks
 - Mixed Review Stoichiometry Answer Key Public Domain eBooks
 - Mixed Review Stoichiometry Answer Key eBook Subscription Services
 - Mixed Review Stoichiometry Answer Key Budget-Friendly Options
6. Navigating Mixed Review Stoichiometry Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Mixed Review Stoichiometry Answer Key Compatibility with Devices
 - Mixed Review Stoichiometry Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mixed Review Stoichiometry Answer Key
 - Highlighting and Note-Taking Mixed Review Stoichiometry Answer Key
 - Interactive Elements Mixed Review Stoichiometry Answer Key
8. Staying Engaged with Mixed Review Stoichiometry Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mixed Review Stoichiometry Answer Key
9. Balancing eBooks and Physical Books Mixed Review Stoichiometry Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mixed Review Stoichiometry Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mixed Review Stoichiometry Answer Key
 - Setting Reading Goals Mixed Review Stoichiometry Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mixed Review Stoichiometry Answer Key
 - Fact-Checking eBook Content of Mixed Review Stoichiometry Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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