

Example:

25. Three grams of carbon and 15 g of oxygen are placed in a closed vessel and heated to 1000°C. What mass of carbon dioxide is formed?

$$25 \text{ g } C + \frac{1 \text{ mol } C}{12 \text{ g } C} = \frac{8}{12} \text{ mol } C = \frac{94 \text{ g } CO_2}{1 \text{ mol } C} = \boxed{74.67 \text{ g } CO_2}$$

26. When 10 g of aluminum reacts with oxygen, 25 g of Al_2O_3 is formed. How much oxygen is required for the reaction?



$$10 \text{ g } Al = \frac{1 \text{ mol } Al}{27 \text{ g } Al} = \frac{2 \text{ mol } Al_2O_3}{2 \times 27 \text{ g } Al} = \frac{102 \text{ g } Al_2O_3}{1 \text{ mol } Al} = 60.67 \text{ g } Al_2O_3 \text{ (theoretical yield)}$$

$$\% \text{ Yield} = \frac{25 \text{ g } Al_2O_3}{60.67 \text{ g } Al_2O_3} \times 100 = \boxed{73.8\%}$$

27. In the reaction of problem 26, what is the limiting reagent when 10 g of Al and 15 g of O_2 react with 10 g of Al?

$$10 \text{ g } Al = \frac{1 \text{ mol } Al}{27 \text{ g } Al} = \frac{2 \text{ mol } Al_2O_3}{2 \times 27 \text{ g } Al} = \frac{2 \text{ g } Al_2O_3}{1 \text{ mol } Al} = 3.0 \text{ g } Al_2O_3$$

$$25 \text{ g } O_2 = \frac{1 \text{ mol } O_2}{32 \text{ g } O_2} = \frac{2 \text{ mol } Al_2O_3}{2 \times 32 \text{ g } O_2} = \frac{2 \text{ g } Al_2O_3}{1 \text{ mol } O_2} = \boxed{2.78 \text{ g } Al_2O_3} \text{ (maximum product)}$$

Limiting Reagent = O_2

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Hussin A.Rothana



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