

Name \_\_\_\_\_

Section \_\_\_\_\_

## Experiment 10

## Advance Study Assignment: Analysis of an Aluminum-Zinc Alloy

1. On the following page, construct a graph of  $N_{H_2}$  vs. % Al. To do this, refer to Equation 11 and the discussion preceding it. Note that a plot of  $N_{H_2}$  vs. % Al should be a straight line (why?). To fix the position of a straight line, it is necessary to locate only two points. The most obvious way to do this is to find  $N_{H_2}$  when % Al = 0 and when % Al = 100. If you wish you may calculate some intermediate points (for example,  $N_{H_2}$  when % Al = 50, or 20, or 80); all these points should lie on the same straight line. To use a spreadsheet, set up Equation 11 for different Al percentages. Graph  $N_{H_2}$  vs. % Al.
2. A student obtained the following data in this experiment. Fill in the blanks in the data and make the indicated calculations:

Mass of gelatin capsule	0.1134 g	Ambient temperature, $t$	23°C
Mass of capsule plus alloy sample	0.3212 g	Ambient temperature, $T$	296 K
Mass of alloy sample, $m$	0.2084 g	Atmospheric pressure	732 mm Hg
Mass of empty beaker	154.3 g	Vapor pressure of $H_2O$ at $t$ (Appendix I)	21.1 mm Hg
Mass of beaker plus displaced water	401.6 g		$732 - 21.1 = 710.9 \text{ mm Hg}$
Mass of displaced water	247.3 g	Pressure of dry $H_2$ , $P_{H_2}$ (Eq. 7)	710.7 mm Hg
Volume of displaced water (density = 1.00 g/mL)	247.3 mL	Pressure of dry $H_2$	0.9354 atm
Volume, $V$ , of $H_2$ = Volume of displaced water	247.3 mL		0.2473 L

Find the number of moles of  $H_2$  evolved,  $n_{H_2}$  (Eq. 5;  $V$  in liters,  $P_{H_2}$  in atm,  $T$  in K,  $R = 0.0821$  liter-atm/mole K).

$$0.935 \times 0.2473 \times 296 = 0.0095 \text{ moles } H_2$$

Find,  $N_{H_2}$ , the number of moles of  $H_2$  per gram of sample ( $n_{H_2}/m$ ).

$$\frac{0.0095}{0.2084} = 0.0456 \text{ moles } H_2/g$$

Find the % Al in the sample from the graph prepared for Problem 1.

Find the % Al in the sample by using Equation 11.

\_\_\_\_\_ % Al

\_\_\_\_\_ % Al006C

# Experiment Advance Study Assignment

**Emil J. Slowinski**



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