

## Stoichiometry III – Theoretical Yield

Teacher

If we know what chemical reaction is occurring, we can use mole ratios to determine the limiting and excess reactants. We can do this by using the starting masses of each reactant to see how much of a product could be produced if all of the reactant was consumed. We then compare the number of moles produced by each reactant, and the smallest value is the **theoretical yield** – the amount of a product produced if all of the limiting reactant is consumed.

### Practice Problems

1. Given the following reaction:  $C_3H_8 + 5O_2 \rightarrow 3CO_2 + 4H_2O$

a) Determine the limiting and excess reagents if you started with 14.8 g of  $C_3H_8$  and 33.44 g of  $O_2$ .

$$\frac{14.8 \text{ g } C_3H_8}{44.09 \text{ g } C_3H_8} \times \frac{1 \text{ mol } C_3H_8}{1 \text{ mol } C_3H_8} \times \frac{4 \text{ mol } H_2O}{1 \text{ mol } C_3H_8} \times \frac{18.02 \text{ g } H_2O}{1 \text{ mol } H_2O} = 24.2 \text{ g } H_2O$$

$$\frac{33.44 \text{ g } O_2}{32.00 \text{ g } O_2} \times \frac{1 \text{ mol } O_2}{1 \text{ mol } O_2} \times \frac{4 \text{ mol } H_2O}{5 \text{ mol } O_2} \times \frac{18.02 \text{ g } H_2O}{1 \text{ mol } H_2O} = 15.1 \text{ g } H_2O$$

Limiting:  $O_2$   
Excess:  $C_3H_8$

b) What is the theoretical yield of  $H_2O$ .  $15.1 \text{ g } H_2O$

2. Given the following reaction:  $Al_2(SO_4)_3 + 6NaOH \rightarrow 3Na_2SO_4 + 2Al(OH)_3$

a) Determine the limiting and excess reagents if 8.1 g of  $Al_2(SO_4)_3$  is reacted with 8.1 g of  $NaOH$

$$\frac{8.1 \text{ g } Al_2(SO_4)_3}{294.14 \text{ g } Al_2(SO_4)_3} \times \frac{1 \text{ mol } Al_2(SO_4)_3}{1 \text{ mol } Al_2(SO_4)_3} \times \frac{2 \text{ mol } Al(OH)_3}{1 \text{ mol } Al_2(SO_4)_3} \times \frac{78.00 \text{ g } Al(OH)_3}{1 \text{ mol } Al(OH)_3} = 4.3 \text{ g } Al(OH)_3$$

$$\frac{8.1 \text{ g } NaOH}{40.00 \text{ g } NaOH} \times \frac{1 \text{ mol } NaOH}{1 \text{ mol } NaOH} \times \frac{2 \text{ mol } Al(OH)_3}{6 \text{ mol } NaOH} \times \frac{78.00 \text{ g } Al(OH)_3}{1 \text{ mol } Al(OH)_3} = 5.3 \text{ g } Al(OH)_3$$

Limiting:  $Al_2(SO_4)_3$   
Excess:  $NaOH$

b) What is the theoretical yield of aluminum hydroxide.  $4.3 \text{ g } Al(OH)_3$

3. Given the following reaction:  $4Al_2O_3 + 9Fe \rightarrow 3Fe_3O_4 + 8Al$

a) Determine the limiting and excess reagents if 25.4 g of  $Al_2O_3$  is reacted with 10.2 g of  $Fe$ .

$$\frac{25.4 \text{ g } Al_2O_3}{101.96 \text{ g } Al_2O_3} \times \frac{1 \text{ mol } Al_2O_3}{1 \text{ mol } Al_2O_3} \times \frac{3 \text{ mol } Fe_3O_4}{4 \text{ mol } Al_2O_3} \times \frac{231.55 \text{ g } Fe_3O_4}{1 \text{ mol } Fe_3O_4} = 43.3 \text{ g } Fe_3O_4$$

$$\frac{10.2 \text{ g } Fe}{55.85 \text{ g } Fe} \times \frac{1 \text{ mol } Fe}{1 \text{ mol } Fe} \times \frac{3 \text{ mol } Fe_3O_4}{9 \text{ mol } Fe} \times \frac{231.55 \text{ g } Fe_3O_4}{1 \text{ mol } Fe_3O_4} = 14.1 \text{ g } Fe_3O_4$$

Limiting:  $Fe$   
Excess:  $Al_2O_3$

b) What is the theoretical yield of iron(II,III) oxide (aka magnetite).

$$14.1 \text{ g } Fe_3O_4$$

# Experiment 3stoichiometry Answers

**Marc le Maire, Raymond Chabaud, Guy  
Hervé**

### **Experiment 3stoichiometry Answers:**

The Word and the World Michael Mulkay, 2024-02-09 First published in 1985 *The Word and the World* is a significant empirically based contribution to the sociological analysis of scientists discourse and scientific culture It also offers a radical departure from established forms of sociological discourse which has far reaching implications for all areas of sociological study The central aim of this book is to widen the range of textual forms used in empirical sociological research and thereby to increase sociologists ability to furnish interesting valid and practically useful conclusions The unusual textual forms employed for the purposes of sociological analysis are the play the investigative inquiry the participant analyst dialogue and the parody One of the main substantive topics that provides the material for these explorations in form is the scientific debate This book will be of interest to students of sociology rhetoric linguistics and science

### **Working with Chemistry**

Donald J. Wink, Sharon Fetzer-Gislason, Julie Ellefson Kuehn, 2004-02-20 With this modular laboratory program students build skills using important chemical concepts and techniques to the point where they are able to design a solution to a scenario drawn from a professional environment The scenarios are drawn from the lives of people who work with chemistry every day ranging from field ecologists to chemical engineers and include many health professionals as well

### **Study Guide to**

**Accompany Basics for Chemistry** Martha Mackin, 2012-12-02 *Study Guide to Accompany Basics for Chemistry* is an 18 chapter text designed to be used with *Basics for Chemistry* textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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**General Chemistry** B. Richard Siebring, Mary Ellen Schaff, 1980 CliffsNotes AP Chemistry Angela Woodward Spangenberg, 2016-01-12 Test prep for the AP Chemistry exam with 100% brand new content that reflects recent exam changes Addressing the major overhaul that the College Board recently made to the AP Chemistry exam this AP Chemistry test prep guide includes completely brand new content tailored to the exam administered every May Features of the guide include review sections of the six big ideas that the new exam focuses on Fundamental building blocks Molecules and interactions Chemical reactions Reaction rates Thermodynamics Chemical equilibrium Every section includes review questions and answers Also included in the guide are two full length practice tests as well as a math review section and sixteen discrete laboratory exercises to prepare AP Chemistry students for the required laboratory experiments section on the exam **Resources in Education** ,1995 **Journal of Chemical Education** ,1926 Includes Report of New England Association of Chemistry Teachers and Proceedings of the Pacific Southwest Association of Chemistry Teachers

**Pp/Chemistry** BarCharts, Inc., 2008-06-18 **Mineral Scale Formation and Inhibition** Z. Amjad, 2013-06-29 This book documents the proceedings of the symposium Mineral Scale Formation and Inhibition held at the American Chemical Society Annual Meeting August 21 to 26 1994 in Washington D C The symposium sponsored by the Division of Colloid and Surface Chemistry was held in honor of Professor George H Nancollas for his pioneering work in the field of crystal growth from solution A total of 30 papers were presented by a wide spectrum of scientists This book also includes papers that were not presented but were in the symposium program The separation of a solid by crystallization is one of the oldest and perhaps the most frequently used operations in chemistry Because of its widespread applicability in recent years there has been considerable interest exhibited by academic and industrial scientists in understanding the mechanisms of crystallization of sparingly soluble salts The salt systems of great interest in industrial water treatment area i e cooling and boiler include

carbonates sulfates phosphates and phosphonates of alkaline earth metals Although not as common as calcium carbonate and calcium sulfate barium and strontium sulfates have long plagued oil field and gas production operations The build up of these sparingly soluble salts on equipment surfaces results in lower heat transfer efficiency increased corrosion rates increased pumping costs etc In the laundry application insoluble calcium carbonate tends to accumulate on washed fabrics and washing equipment parts resulting in undesirable fabric encrustation or scaling

*Chemistry, Study Guide* Bernice G. Segal, Peter S. Shenkin, 1989-02-14 This Second Edition of the first year chemistry text known for its clarity of exposition and its large number of illustrative worked problems contains a more rigorous treatment of electrochemistry chemical equilibrium and thermochemistry Worked examples now number over 300 and exercises over 1460

**Solutions Manual to Accompany Physical Chemistry for the Life Sciences** C. A. Trapp, M. P. Cady, 2011 This solutions manual contains fully worked solutions to all end of chapter discussion questions and exercises featured in *Physical Chemistry for the Life Sciences*

**Catalog** Community College of Denver, 2009

**Foundations of Inorganic Chemistry** Gary Wulfsberg, 2017-11-02

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CLASS 11 CHEMISTRY 5 SOLVED "CASE STUDIES" Priti Singhal,2024-11-17 This book is structured to align with the latest syllabus and curriculum guidelines ensuring that the content is both relevant and rigorous Each chapter begins with a clear set of learning objectives providing a roadmap for students to understand what they will achieve by the end of the chapter We have included numerous diagrams illustrations and real life examples to make complex concepts more accessible and engaging

**Materials Aspect of Thermoelectricity** Ctirad Uher,2016-11-25 In recent years novel families of materials have been discovered and significant improvements in classical thermoelectric materials have been made Thermoelectric generators are now being used to harvest industrial heat waste and convert it into electricity This is being utilized in communal incinerators large smelters and cement plants Leading car and truck companies are developing thermoelectric power generators to collect heat from the exhaust systems of gasoline and diesel engines Additionally thermoelectric coolers are being used in a variety of picnic boxes vessels used to transport transplant organs and in air conditioned seats of mid size cars Consisting of twenty one chapters written by top researchers in the field this book explores the major advancements being made in the material aspects of thermoelectricity and provides a critical assessment in regards to the broadening of application opportunities for thermoelectric energy conversion

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*Modern Experimental Chemistry* George W. Jr. Latimer, 2012-12-02 Modern Experimental Chemistry provides techniques of qualitative analysis that reinforce experiments on ionic equilibria This book includes the determination of water in hydrated salts identification of an organic compound after determining its molecular weight and nonaqueous titration of a salt of a weak acid The calculation of chemical stoichiometry calculation of thermodynamic properties by determining the change in equilibrium with temperature and chromium chemistry are also covered This compilation contains enough experiments for classes which have six hours of laboratory two 3 hour meetings per week to last two semesters This publication is intended for chemistry students as an introductory manual to chemistry laboratory

**5 Steps to a 5: AP Chemistry 2018 Elite Student Edition** Mary Millhollon, Richard H. Langley, 2017-07-28 Get ready to ace your AP Chemistry Exam with this easy to follow multi platform study guide 5 Steps to a 5 AP Chemistry 2018 Elite Student Edition introduces an effective 5 step study plan to help you build the skills knowledge and test taking confidence you need to achieve a high score on the exam This popular test prep guide matches the latest course syllabus and latest exam You ll get online help four full length practice tests two in the book and two online detailed answers to each question study tips and important information on how the exam is scored Because this guide is accessible in print and digital formats you can study online via your mobile device straight from the book or any combination of the three With the new 5 Minutes to a 5 section you ll also get an extra AP curriculum activity for each school day to help reinforce the most important AP concepts With only 5 minutes a day you can dramatically increase your score on exam day 5 Steps to a 5 AP Chemistry 2018 Elite Student Edition features New 5 Minutes to a 5 Concise activities reinforcing the most important AP concepts and presented in a day to day study format Access to the entire Cross Platform Prep Course in Chemistry 4 Practice Exams 2 in the book 2 online Powerful analytics you can use to assess your test readiness Flashcards games social media support and more

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<https://staging.conocer.cide.edu/About/virtual-library/fetch.php/Harold%20The%20Last%20Of%20The%20Saxon%20Kings%20Of%20Falkl.pdf>

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