

Study Guide – Unit 1 Periodic Table **ANSWERS**

Structure of an atom

- List the three subatomic particles which make up an atom. Include their location, charge, and atomic mass.
 - Electron** Negative charge Electron Cloud
 - Proton** Positive charge Nucleus
 - Neutron** Neutral charge Nucleus
- In a neutral atom, which two subatomic particles are equal in number to each other?
Protons and Electrons
- What does the atomic number of an element indicate?
The number of protons in the atom of that element
- What is an isotope?
Atoms with the same number of protons, but different masses due to a different number of neutrons
- Will two different atoms of carbon always have the same number of protons?
Yes

Element key

13
Al
26.982
Aluminum

- What are the number of protons, neutrons, and electrons in an atom of aluminum?
 - Protons = 13**
 - Neutrons = 14**
 - Electrons = 13**

Bohr Model

- How do you know how many energy levels (orbits) a Bohr model will have?
 - By the period the element is located in**
- How many electrons can each energy level hold?
1st – 2 2nd – 8 3rd – 18

The Periodic Table

- How is the periodic table arranged? **Increasing atomic number**
- What are the vertical columns \updownarrow called? **Groups or Families**
- What are the horizontal rows \leftrightarrow called? **Periods**
- Why are elements put into groups? **Same number of valence electrons**
- What is a valence electron? **Number of electrons on the outer energy level (shell)**
- Know how many valence electrons each group has
- Which group of metals are highly reactive and why? **Alkali Metals. Only 1 ve⁻ to lose**
- Which group of non-metals are highly reactive and why? **Halogens. Only 1 ve⁻ to gain**
- Name some physical properties of metals: **luster, ductile, malleable, conductor of heat and electricity**
- Name some physical properties of non metals: **no luster, brittle, not a good conductor**
- What are metalloids and where are they found on the periodic table? **Share properties of metals and metalloids. On zig-zag line starting at Boron.**

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Yi-Tong Ma



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