



# Excalibur Electronic User Manual

**Clemens Wendtner**



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web log H<sub>10</sub> 4.05891 × 10<sup>-5</sup> mol dm<sup>-3</sup> log pK<sub>a</sub> pH 3.7540503010030050 and HCOOH 0.1 so HCOO<sup>-</sup> 0.200 mol dm<sup>-3</sup> amount vol x

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web only the better candidates could write successive ionisation equations in part c while part d about the pH of equimolar strong and weak acid solutions was poorly answered

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web a substance may be assigned to one of our four conceivable categories it may be an acid or a base but in addition it may be both an acid and a base or it may be neither an acid nor

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web to detect the presence of acids and bases Michael adds a few drops of phenolphthalein indicator to two colorless solutions a and b he observes the following changes in the

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web section 2 pH of strong acids number of protons released monoprotic acid acid that releases one H<sup>+</sup> ion per molecule e.g.

hcl hydrochloric acid hno<sub>3</sub> nitric acid ch

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web problem 1 which of the following are not acids ch<sub>3</sub>cooh co<sub>2</sub> hno<sub>2</sub> hcooh cc<sub>4</sub> c h<sub>3</sub> c o o h c o<sub>2</sub> h n o<sub>2</sub> h c o o h c c l<sub>4</sub>

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