

MODEL QUESTION PAPER-2 : 2020-21
CHEMISTRY (THEORY)

MM : 70

Time : 3 Hours

General Instructions. Read the following instructions carefully.

- (a) There are 33 questions in this question paper. All questions are compulsory.
- (b) Section A: Q. No. 1 to 2 are case-based questions having four MCQs or Reason Assertion type based on given passage each carrying 1 mark.
- (c) Section A: Question 3 to 16 are MCQs and Reason Assertion type questions carrying 1 mark each.
- (d) Section B: Q. No. 17 to 25 are short answer questions and carry 2 marks each.
- (e) Section C: Q. No. 26 to 30 are short answer questions and carry 3 marks each.
- (f) Section D: Q. No. 31 to 33 are long answer questions carrying 5 marks each.
- (g) There is no overall choice. However, internal choices have been provided.
- (h) Use of calculators and log tables is not permitted.

Section-A (Objective Type)**Objective Type Questions**

1. Read the passage given below and answer the following questions: (1×4=4)

The d-block of the periodic table contains the elements of the groups 3 – 10 and are known as transition elements. In general, the electronic configuration of these elements is $(n-1) d^{1-10} ns^2$. The *d*-orbitals of the penultimate energy level in their upper receive electrons giving rise to the three rows of the transition metals i.e., 3d, 4d and 5d series. However, Zn, Cd and Hg are not regarded as transition elements. Transition elements exhibit certain characteristic properties like variable oxidation states, complex formation, formation of coloured ions and alloys, catalytic activity, etc. Transition metals are hard (except Zn, Cd and Hg) and have a high melting point.

The following questions are multiple choice questions. Choose the most appropriate answer:

- (i) Electronic configuration of a transition element X in +3 oxidation state is $[Ar]3d^5$. What is its atomic number?
- (a) 25 (b) 26 (c) 27 (d) 24
- (ii) The electronic configuration of $Cu(H)$ is $3d^9$ whereas that of $CuCl$ is $3d^{10}$. Which of the following is correct?
- (a) $Cu(H)$ is more stable
- (b) $Cu(H)$ is less stable
- (c) $Cu(H)$ and $CuCl$ are equally stable
- (d) Stability of $Cu(H)$ and $CuCl$ depends on nature of copper salt

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