

Electron Configuration Practice
Chemistry

Name: _____

Due Date: _____

How to write an electron configuration:

- Determine the total number of electrons to be represented.
- Use the Aufbau principle to fill the orbitals with electrons for elements 1-23. Refer to electron configuration periodic table for elements after 23
- The sum of the superscripts should equal the total number of electrons. Example:
 $12\text{Mg } 1s^2 2s^2 2p^6 3s^2$

I. Configuration Writing Practice

- Write a ground state electron configuration for each neutral atom. Ground state means that all of the lowest possible energy levels (up to the proper number of electrons for the element) are filled.

1. Na: $1s^2 2s^2 2p^6 3s^1$

2. Pb: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^2$

3. Sr: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2$

4. U: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 5f^3 6d^1$

5. N: $1s^2 2s^2 2p^3$

* 6. Ag: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1 4d^{10}$

7. Ti: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^2$

8. Ce: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^1 5d^1$

9. Cl: $1s^2 2s^2 2p^6 3s^2 3p^5$

10. Hg: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10}$

Electron Configurations Packet With Answers

CH Cherryholmes



Electron Configurations Packet With Answers:

The book delves into Electron Configurations Packet With Answers. Electron Configurations Packet With Answers is an essential topic that needs to be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Electron Configurations Packet With Answers, encompassing both the fundamentals and more intricate discussions.

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- Chapter 1: Introduction to Electron Configurations Packet With Answers
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- Chapter 3: Electron Configurations Packet With Answers in Everyday Life
- Chapter 4: Electron Configurations Packet With Answers in Specific Contexts
- Chapter 5: Conclusion

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