

CHAPTER 13 REVIEW ACTIVITY

Text Reference: Section 13-10

Writing Electron Configurations

The filling order for electrons in energy sublevels is:

1s, 2s, 2p, 3s, 3p, 4s, 3d, 4p, 5s, 4d,
5p, 6s, 4f, 5d, 6p, 7s, 5f, 6d, 7p

Each *s* sublevel contains 1 orbital; each *p* contains 3; each *d* contains 5; and each *f* contains 7. Each orbital can contain at most 2 electrons.

An electron configuration can be written out by assigning electrons to the sublevels in the order listed, until the number of electrons assigned equals the atomic number (*N*) of the atom.

Examples

Determine the configurations of these elements.

Beryllium (*N* = 4)

$1s^2 2s^2$

Aluminum (*N* = 13)

$1s^2 2s^2 2p^6 3s^2 3p^1$

Bromine (*N* = 35)

$1s^2 2s^2 2p^6 3s^2 3p^4 4s^2 3d^{10} 4p^5$

Write the electron configuration for each of the following elements.

1. Calcium (*N* = 20)

1. _____

2. Lithium (*N* = 3)

2. _____

3. Argon (*N* = 18)

3. _____

4. Iron (*N* = 26)

4. _____

5. Sodium (*N* = 11)

5. _____

6. Oxygen (*N* = 8)

6. _____

7. Iodine (*N* = 53)

7. _____

8. Dysprosium (*N* = 66)

8. _____

9. Radium (*N* = 88)

9. _____

10. Fermium (*N* = 100)

10. _____

Electron Configuration Chem Worksheet 5 6 Answers

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