

8 • Electron Configurations & Periodicity

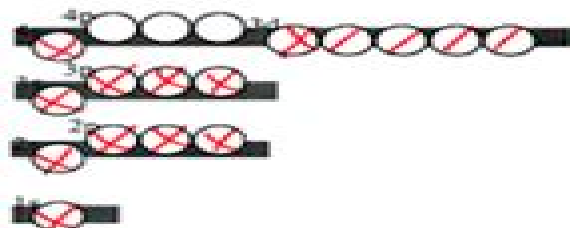
WRITING ELECTRON CONFIGURATIONS

For each given element, fill in the orbital diagram and then write the electron configuration for the element.

1.	2.	3.	4.	5.	6.
Element: Ar # of e ⁻ 's: <u>18</u>	Element: Mg # of e ⁻ 's: <u>12</u>	Element: N # of e ⁻ 's: <u>7</u>	Element: Li # of e ⁻ 's: <u>3</u>	Element: P # of e ⁻ 's: <u>15</u>	Element: Cl # of e ⁻ 's: <u>17</u>

Write the electron configurations of each of these in **long form** and **short form**:

- Ar 1s² 2s² 2p⁶ 3s² 3p⁶
 Ar [Ne] 3s² 3p⁶
- Mg 1s² 2s² 2p⁶ 3s²
 Mg [Ne] 3s²
- N 1s² 2s² 2p³
 N [He] 2s² 2p³
- Li 1s² 2s¹
 Li [He] 2s¹
- P 1s² 2s² 2p⁶ 3s² 3p³
 P [Ne] 3s² 3p³
- Cl 1s² 2s² 2p⁶ 3s² 3p⁵
 Cl [Ne] 3s² 3p⁵



7. Fill in the orbital diagram for the element, Fe, and write the electron configuration of Fe in the long and short form.



Electron Configuration And Periodicity Lab Answers

Shasha Hu



Electron Configuration And Periodicity Lab Answers:

The Software Encyclopedia 2001 ,2001 High School Chemistry Unlocked The Princeton Review,2016-11-29
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