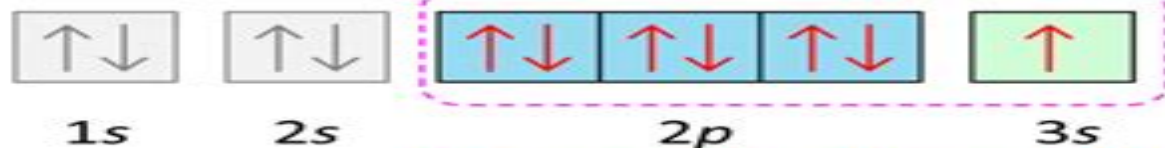


Electron Configuration of Main Group Cations

Electron configuration

Na $1s^2 2s^2 2p^6 3s^1$

Orbital diagram

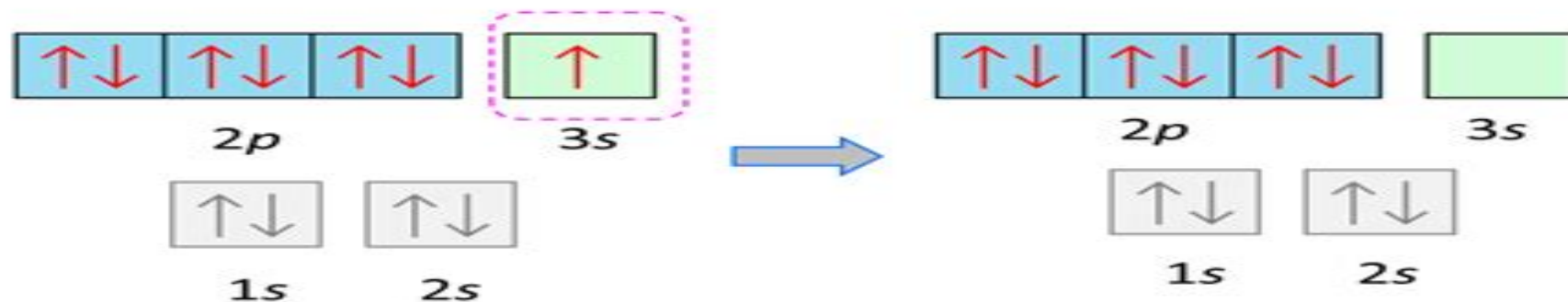


Focus on the valence shell

The charge is +1, therefore, one electron is lost:



The electron must be from the outermost valence orbital – $3s^1$



Na $1s^2 2s^2 2p^6 3s^1$

Na⁺ $1s^2 2s^2 2p^6$

↓ Same as

Ne $1s^2 2s^2 2p^6$



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Cations of main group metals attain the electron configuration of the noble gas that precedes them in the periodic table. They are **isoelectronic** with the previous noble gas.

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