

Chapter 17: Electrochemistry Study Guide

Electrochemistry-

- The study of the interchange of chemical and electrical energy
- Processes are oxidation-reduction reactions in which:
 - The energy released by a spontaneous reaction is converted to electricity
 - The electrical energy is used to cause a nonspontaneous reaction
- Galvanic cell- chemical energy is transformed to electrical energy through red-ox
- Electrolytic cell- electrical energy is used to produce a chemical change

Oxidation Number:

- Free element that have an oxidation number of zero- also include BrINClHOF
- Monoatomic ions have oxidation numbers that are equal to the charge of their ions
- Oxygen's number= -2 except when O_2^{2-} , then it's -1
- Hydrogen's number= +1, except when with binary compounds, then it is -1
- Group 1= +1, Group 2=+2
- Fluorine is always -1
- Sum of the oxidation numbers should be equal to the charge.

Oxidation: $2 \text{Mg} \rightarrow 2\text{Mg}^{2+} + 4\text{e}^-$

Reduction: $\text{O}_2 + 4\text{e}^- \rightarrow 2\text{O}^{2-}$

"An-Ox Fat-Red-Cat"

Galvanic Cells:

- Anode- Oxidation Cathode- Reduction
- E° - the driving force behind the electron transfer
 - Measured in volts (V)
 - Defined as joules of work per coulomb of charge
 - $E^\circ = \frac{-w}{q}$
- Always flows from anode to cathode
- The salt bridge allows the charged particles to flow back and forth to keep it flowing
- Electrons always go towards the one with the most positive concentration
- The difference in electrical potential between the anode and cathode is called cell voltage, electromotive force, or cell potential

Standard Reduction Potential (E°)- the voltage associated with a reduction reaction at an electrode when all solutes are 1M and all gases are at 1atm

- $E^\circ = \text{Cathode} - \text{Anode} = \text{Reduction} - \text{Oxidation}$
- To find E° , flip the smaller Cell potential value and subtract the E° value
- In a galvanic cell, it must have a positive E° in order to go.

Electrochemistry Study Guide Answers

Clemens Wendtner



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