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Manual Proton Saga

H Kauffman



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web 1 define each remember oil rig oxidation is loss of e reduction is gain of e a oxidation b reduction c oxidizing agent d reducing agent 2 write half reactions for each of the following atoms or ions label each as oxidation or reduction use the activity series chart a al b ba2 c br 2 d ca e ga3 f h 2 g h

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web describe redox reactions in terms of electron transfer recall that oxidation involves loss of electrons and reduction involves gain of electrons represent simple redox reactions using half equations identify oxidising and reducing agents for simple redox reactions understand the reactivity of metals and metal displacement reaction

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you $23 \text{ Fe}^{2+} + \text{Co} \rightarrow \text{Co}^{2+} + \text{Fe}$
oxidation $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^{-}$ reduction $\text{Fe}^{2+} + 2\text{e}^{-} \rightarrow \text{Fe}$ $24 \text{ 3 Ag} + \text{Ni} \rightarrow \text{Ni}^{3+} + 3 \text{ Ag}$ oxidation reduction $25 \text{ Cu}^{2+} + \text{Pb} \rightarrow \text{Pb}^{2+} + \text{Cu}$ oxidation reduction 26 O_2

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s0 to s2 reduced ox ag 3 2cs br2