Digities

Chemical Quantities



THE MOLE AND QUANTIFYING MATTER

10.1 The Mole: A Measurement of Matter

For students using the Foundation edition, assign problems 1–15.

Essential Understanding The mole represents a large number of very small particles.



Frayer Model The Frayer Model is a vocabulary development tool. The center of the diagram shows the concept being defined, while the quadrants around the concept are used for providing the details. Use this model when you want to understand a vocabulary term in more detail.

As you read Lesson 10.1, use the Frayer Model below. Complete the diagram for the term mole.

Definition in your own words	Facts/characteristics
A mole is 6.02 × 10 ²¹ of anything.	A mole contains Avogadro's number of anything. The mass of a mole of an element is the atomic mass of the element in grams.
How moles are used Conversion factors can be	Examples
used to change the number of particles to the number of moles or	1 mol Na = 23.0 g Na = 6.02 × 10 ²³ Na atoms.
the number of moles to the number of particles.	1 mol $H_2O = 18.0 \text{ g } H_2O$ = $6.02 \times 10^{43} H_2O$ molecules.

Lesson 6 Reading And Study Workbook Copyright

Rachel Sandford

Lesson 6 Reading And Study Workbook Copyright:

This book delves into Lesson 6 Reading And Study Workbook Copyright. Lesson 6 Reading And Study Workbook Copyright is a crucial topic that must be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Lesson 6 Reading And Study Workbook Copyright, encompassing both the fundamentals and more intricate discussions.

- 1. The book is structured into several chapters, namely:
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