

## অধ্যায় ১: ল্যাবরেটরির নিরাপদ ব্যবহার



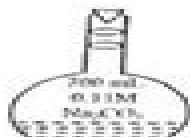
পরীক্ষার কপূর, এখানে কোর পরীক্ষা, আরও কলেজের শিক্ষার্থীদের কলেজসমূহের পরীক্ষা এবং মডেল টেস্টের প্রস্তুতি পূরণের উত্তর অধ্যয়নিক সেবা করে। এগুলোর অনুষ্ঠান করতে তুমি এ অধ্যয়নিক থেকে সেবারের সুজনীন প্রস্তুতি

এদের উত্তর নিচে পড়বে সহজেই।

প্রশ্ন



পাত্রে-A



পাত্রে-B

- ক. মোলার প্রবণ কী?  
 খ. "কোনো পানিদ্রুত দ্রবণের দ্রাব্যতা, তাপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়"— ব্যাখ্যা করো।  
 গ. A-পাত্রে বিদ্যমান এলিডের ভর কত নির্ণয় করো।  
 ঘ. দুটি মিশ্রিত করলে, মিশ্রণটি কী পরিমাণ অম্লীয় বা ক্ষারীয় হবে— বিশ্লেষণ করো।

১ নং প্রশ্নের উত্তর

1000 mL দ্রবণে কোন দ্রবের 1 mol দ্রবীভূত থাকলে উক্ত দ্রবকে মোলার দ্রবণ বলে।

কোনো পানিদ্রুত দ্রবণের দ্রাব্যতা তাপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়। এর কারণ, তাপমাত্রা বৃদ্ধির সাথে সাথে দ্রাব্যতা বৃদ্ধি পেতে থাকলেও একটি নির্দিষ্ট তাপমাত্রার কোনো পানি দ্রুত দ্রবণ নিষ্কাশিত হয়। তখন দ্রাব্যতা হ্রাস পেতে থাকে। উদাহরণস্বরূপ  $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$  এর কথা বলা যায়।  $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$  এর ক্ষেত্রে পানিতে দ্রবীভূত করার সময় বিয়োজন প্রক্রিয়াটি তাপস্বরী বিক্রিয়া কিন্তু  $32.4^\circ\text{C}$  তাপমাত্রার উপরে  $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$  যখন  $\text{Na}_2\text{SO}_4$  এ পরিণত হয় তখন বিয়োজন প্রক্রিয়াটি তাপোৎপাদী হয়। এ কারণে  $32.4^\circ\text{C}$  তাপমাত্রার উপরে তাপমাত্রা বৃদ্ধিজনিত কারণে দ্রাব্যতার হ্রাস ঘটে।

আমরা জানি,  
 $w = SMV$   
 $= (0.5 \times 98 \times 0.25) \text{ g}$   
 $= 12.25 \text{ g}$   
 অতএব, দ্রবীভূত  $\text{H}_2\text{SO}_4$  এর ভর = 12.25 g

উদীয়কের A-পাত্রে অবস্থিত,  $\text{H}_2\text{SO}_4$  এর আণবিক ভর,  
 $M = 1 \times 2 + 32 + 16 \times 4$   
 $= 2 + 32 + 64$   
 $= 98 \text{ g/mol}$   
 দ্রবণের ঘনমাত্রা,  
 $S = 0.5 \text{ M}$   
 দ্রবণের আয়তন,  
 $V = 250 \text{ mL}$   
 $= \frac{250}{1000} = 0.25 \text{ L}$   
 দ্রবীভূত  $\text{H}_2\text{SO}_4$  এর ভর,  $w = ?$

পাত্রে-A:  $\text{H}_2\text{SO}_4$  এর মোল সংখ্যা =  $\left(0.5 \times \frac{250}{1000}\right) \text{ mol}$   
 $= 0.125 \text{ mol}$

পাত্রে-B:  $\text{Na}_2\text{CO}_3$  এর মোল সংখ্যা =  $\left(0.11 \times \frac{700}{1000}\right) \text{ mol}$   
 $= 0.077 \text{ mol}$

সংঘটিত বিক্রিয়া:  $\text{H}_2\text{SO}_4 + \text{Na}_2\text{CO}_3 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{CO}_2$   
 সমীকরণ হতে পাই, 1 mol  $\text{H}_2\text{SO}_4$  = 1 mol  $\text{Na}_2\text{CO}_3$   
 1 mol  $\text{Na}_2\text{CO}_3$  বিক্রিয়া করে  $\text{H}_2\text{SO}_4$  এর 1 mol এর সাথে 0.077 mol

অবশিষ্ট  $\text{H}_2\text{SO}_4$  এর পরিমাণ =  $(0.125 - 0.077) \text{ mol}$   
 $= 0.048 \text{ mol}$

পাত্রে-A এর দ্রবণের সাথে পাত্রে-B এর দ্রবণের বিক্রিয়ার ফলে 0.048 mol  $\text{H}_2\text{SO}_4$  অবশিষ্ট থাকায় মিশ্রিত দ্রবণের প্রকৃতি হবে অম্লীয়।

একদল ছাত্র ল্যাবরেটরিতে X-পদ্ধতিতে 10 mg এর ভরে ও পরিমাণ এবং Y-পদ্ধতিতে 0.5 g পরিমাণ 2%  $\frac{W}{V}$  HCl এর নমুনা তৈরি করে।

- ক. লোভ কাকে বলে?  
 খ.  $\text{Cl}_2$  অপোলার— ব্যাখ্যা করো।  
 গ. উদীয়কের এলিডটির মোলার ঘনমাত্রা নির্ণয় করো।  
 ঘ. উদীয়কের কোন পদ্ধতি পরিবেশবান্ধব? বিশ্লেষণ করো।

২ নং প্রশ্নের উত্তর

নিউক্লিয়াসের চারদিকে যে নির্দিষ্ট ত্রিমাত্রিক স্থানে কোন ইলেক্ট্রন পাওয়ার সম্ভাব্যতা শূন্য তাকে লোভ বলে।

সমযোজী বন্ধনে আকর্ষণ উভয় পরমাণুর তড়িৎ ঋণাত্মকতা কম হলে অর্থাৎ তড়িৎ ঋণাত্মকতার পার্থক্য শূন্য হলে সমযোজী বন্ধন বিশুদ্ধ ও অপোলার হয়। আর তড়িৎ ঋণাত্মকতার পার্থক্য 0.5 এর কম হলে অপোলার হয়। যেহেতু Cl এর তড়িৎ ঋণাত্মকতা 3.0 ও C অপূর্তে তড়িৎ ঋণাত্মকতার পার্থক্য  $(3 - 3) = 0$ , তাই  $\text{Cl}_2$  বিশুদ্ধ সমযোজী ও অপোলার যৌগ। উল্লেখ্য যৌগিক অণুর ক্ষেত্রে এরূপ থাকে।

উদীয়কে বণিত এলিডটি হলো 2%  $\left(\frac{W}{V}\right)$  HCl।  
 2%  $\left(\frac{W}{V}\right)$  HCl এর অর্থ হলো—  
 100 mL দ্রবণে HCl এর ভর 2g  
 $\therefore 1000 \text{ mL " " " " } \left(\frac{2 \times 1000}{100}\right) \text{ g}$   
 $= 20 \text{ g}$

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**G Psacharopoulos**



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