Name _____

Stoichiometry Worksheet 2: Percent Yield

For each of the problems below:

- a. Write the balanced chemical equation
- b. Identify the given (with units) and what you want to find (with units)
- c. Show set up with units. Check sig figs, give final answer with units and label.
- Using the Hoffman apparatus for electrolysis, a chemist decomposes 36 g of water into its gaseous elements. How many grams of hydrogen gas should she get (theoretical yield)?

Equation:	2 H ₂ O ₍₀₎ →	2 H _{2(c)} +	1 O2 (a)
Before	2 mol	0 mol	0 mol
Change	-2 mol	+2 mol	
After	0 mol	2 mol	

$$36 \text{ g } H_2Ox \frac{1 \text{ mol } H_2O}{18.0 \text{ g } H_2O} = 2.0 \text{ mol } H_2O \times \frac{2 \text{ mol } H_2}{2 \text{ mol } H_2O} = 2 \text{ mol } H_2 \times \frac{2.0 \text{ g } H_2}{1 \text{ mol } H_2} = 4.0 \text{ g } H_2$$

2. Recall that liquid sodium reacts with chlorine gas to produce sodium chloride. You want to produce 581 g of sodium chloride. How many grams of sodium are needed?

Equation:	2 Na ₍₀₎ +	1 Cl _{2(a)} \rightarrow	2 NaCl _{tet}	
Before	9.94 mol	xs mol	0 mool	
Change	-9.94 mol		+9.94 mol	
After	0 mol		9.94 mod	

$$581 \text{ g NaClx} \frac{1 \text{ mol NaCl}}{58.5 \text{ g NaCl}} = 9.93 \text{ mol NaCl} \times \frac{2 \text{ mol Na}}{2 \text{ mol NaCl}} = 9.93 \text{ mol Na} \times \frac{23.0 \text{ g Na}}{1 \text{ mol Na}} = 228 \text{ g Na}$$

3. You eat 180.0 g of glucose (90 M&Ms). If glucose, C6H12O6, reacts with oxygen gas to produce carbon dioxide and water, how many grams of oxygen will you have to breathe in to burn the glucose?

Equation:	1 C6H12O6(a) +	6 O _{2(g)} → 6 CO _{2(g)} + 6 H ₂	$O_{(ab)}$
Before	1.0 mod	xs mol 0 mol 0 mol	
Change	-1.0 mol	-1.0 mol	
After	0 mol	xs mol	

$$180.0 \text{ g C}_{a}\text{H}_{12}\text{O}_{a} \times \frac{1 \text{ mol C}_{a}\text{H}_{12}\text{O}_{a}}{180 \text{ g C}_{a}\text{H}_{12}\text{O}_{a}} = 1.00 \text{ mol C}_{a}\text{H}_{12}\text{O}_{a} \times \frac{6 \text{ mol O}_{2}}{1 \text{ mol C}_{a}\text{H}_{12}\text{O}_{a}} = 6.00 \text{ mol O}_{2}$$

$$6.0 \text{ mol } O_2 x \frac{32.0 \text{ g } O_2}{1 \text{ mol } O_2} = 192.0 \text{ g } O_2$$

Guide Chemistry Stoichiometry Answer Key

VM Jensen

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