

## Reaction Kinetics Lab Reporting Worksheet

In science, reporting what has been done in a laboratory setting is incredibly important for communicating, replicating, and validating findings. However, writing scientific reports can be a little overwhelming. There are a set of agreed upon components that the scientific community requires when reporting scientific research. Answer the following questions to describe what occurred during the lab you conducted in Labster. Be sure to use complete sentences and descriptions that fully represent what you experienced. Writing a lab report is less about being correct or incorrect, than it is accurately reporting what happened and why. So, do not worry about reporting data that might seem counterintuitive or unexpected. Focus on clearly communicating what you did and what you observed.

Write your answers on a new line.

- What was the title of the lab you completed?  
Energy surfaces and spontaneous reactions was the lab title.

### 1. Topic

- What was the subject you were trying to understand better in the lab?  
We were trying to understand reaction kinetics and what parameters influence reaction rates. We learned about how temperature changes affect levels of energy and reactions to different molecules. The lab topics are how temperature and concentration affect the rate of the reaction. The optimal reaction to get the drone across the lake is needed.

### 2. Background Information

- What information from the textbook and classroom is relevant for the subject you were trying to gain a better understanding of in the lab? Identify the concepts and explain how they are related to the lab topic.

Decomposition of hydrogen peroxide Hydrogen peroxide ( $\text{H}_2\text{O}_2$ ) is a strong oxidizer and is part of a peroxide group. These very unstable peroxide bonds cause the compound to decompose easily. At elevated temperatures, or with a presence of a catalyst, purified hydrogen peroxide decomposes into oxygen and water. When this decomposition is fast enough, hydrogen peroxide can be used as a monopropellant for submarines and satellites. Reaction rate- is the change in concentration of a reactant or product per unit of time. The following mathematical formula can be used to express the reaction rate:  $\text{Rate} = \Delta[X] / \Delta t$  Which means Reaction rate = (concentration of X at  $t_2$  - concentration of X at  $t_1$ ) / ( $t_2 - t_1$ ) =  $\Delta[X] / \Delta t$

- During the lab, what information from the theory section provided additional background information about the subject? (To review the theory section, launch the lab and click the Theory tab on the top of the data pad). Identify the concepts and explain how they are related to the lab topic.

Collision models the collision model or collision theory is based on kinetic-molecular theory. It explains the effect of concentration and temperature on the reaction rate. According to the model, reactant molecules only react when they collide with the correct orientation and enough energy to overcome the activation energy. Increasing the rate of successful collision increases the reaction rates. Two requirements must be complete for reactants to collide successfully: 1) the collision energy must be equal to or exceed the activation energy. 2) The orientation must allow the formation of new chemical bonds. the activation energy ( $E_a$ ) represents the minimum energy required to initiate a reaction. The reaction rate highly depends on the activation energy. The lower the activation energy, the higher the reaction rate. The enthalpy changes of the reaction,  $\Delta H$ , is estimated as the difference in "heat energy" between the reactants and products (at constant pressure). The molecule at the top of the curve or energy barrier is called the activated complex or the transition state.

# Growth Kinetics Lab Report

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