

## Quarter 3 Exam Review

## Multiple Choice

Always, the choice that best completes the statement or answers the question.

- \_\_\_\_\_ 1. In a chemical reaction
- the mass of the reactants equals the mass of the products.
  - the mass of the product is greater than the mass of reactants.
  - the material atoms in the reactants and products must change.
  - energy is lost must be added to the reactants.
- \_\_\_\_\_ 2. When a ratio, provided by a chemical reaction, compares facts the solution it is called
- a percent.
  - a percent.
  - a percent.
  - the mass of the product.
- \_\_\_\_\_ 3. In writing a chemical equation that produces hydrogen gas, the correct representation of hydrogen gas is
- H
  - H<sub>2</sub>
  - H<sub>2</sub>
  - H<sub>2</sub>
- \_\_\_\_\_ 4. To balance a chemical equation, it may be necessary to adjust the
- coefficients.
  - formulas of the products.
  - subscripts.
  - reactant of products.
- \_\_\_\_\_ 5. According to the law of conservation of mass, the total mass of the reacting substances is
- always more than the total mass of the products.
  - always less than the total mass of the products.
  - sometimes more and sometimes less than the total mass of the products.
  - always equal to the total mass of the products.
- \_\_\_\_\_ 6. In the chemical equation  $2\text{Mg}(s) + \text{O}_2(g) \rightarrow 2\text{MgO}(s)$
- Mg represents the product magnesium.
  - O<sub>2</sub> represents the reactant oxygen.
  - the reaction yield magnesium.
  - O<sub>2</sub> represents the product oxygen gas.
- \_\_\_\_\_ 7. When the reaction  $2\text{FeCl}_3 + 3\text{H}_2 \rightarrow 2\text{Fe} + 6\text{HCl}$  is correctly balanced, what is the coefficient of Fe?
- 4
  - 2
  - 6
  - 3
- \_\_\_\_\_ 8. Which coefficients correctly balance the formula equation  $\text{CaCl}_2 + \text{H}_2\text{O} \rightarrow \text{Ca}(\text{OH})_2$ ?
- 1, 1, 2
  - 1, 2, 1
  - 1, 1, 1
  - 1, 2, 1
- \_\_\_\_\_ 9. In which kind of reaction do two or more substances combine to form a new compound?
- decomposition reaction
  - double displacement reaction
  - single displacement reaction
  - synthesis reaction
- \_\_\_\_\_ 10. The equation  $\text{A} + \text{B} \rightarrow \text{C} + \text{D}$  is the general equation for a
- synthesis reaction.
  - decomposition reaction.
  - single displacement reaction.
  - double displacement reaction.

# General Chemistry Review Exam Guide Answers Packet

**Michael Hoelscher**



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