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## CHAPTER 8

### Electron Configurations and Periodicity

#### ■ SOLUTIONS TO EXERCISES

*Note on significant figures:* If the final answer to a solution needs to be rounded off, it is given first with one nonsignificant figure, and the last significant figure is underlined. The final answer is then rounded to the correct number of significant figures. In multistep problems, intermediate answers are given with at least one nonsignificant figure; however, only the final answer has been rounded off.

- 8.1.
- a. Possible orbital diagram.
  - b. Possible orbital diagram.
  - c. Impossible orbital diagram; there are two electrons in a  $2p$  orbital with the same spin.
  - d. Possible electron configuration.
  - e. Impossible electron configuration; only two electrons are allowed in an  $s$  subshell.
  - f. Impossible electron configuration; only six electrons are allowed in a  $p$  subshell.

# General Chemistry Lab Manual Ebbing Answers

**JE Gale**



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