CLASSIFYING CHEMICAL REACTIONS

14. SO₃ + H₂O → H₂SO₄



Classify the following reactions as synthesis, decomposition, single replacement or double replacement. (AACCI = RC)

repr	acement.	ANSW	EKS/
1.	2KCIO ₃ → 2KCI + 3O ₂	Decon	nposition
2.	HCI + NaOH → NaCI + H ₂ O	Double	Replacement
з.	Mg + 2HCl → MgCl ₂ + H ₂	Single	Replacement.
4.	2H ₂ + O ₂ → 2H ₂ O	Synthesis	
5.	2AI + 3NIBr₂ → 2AIBr₃ + 3NI	Single	Replacement
6.	4AI + 3O₂ → 2AI₂O₃	Synth	ne Sis
7.	2NaCl → 2Na + Cl ₂	Decom	position
8.	$CaCl_2 + F_2 \rightarrow CaF_2 + Cl_2$	Single	Replacement
9.	AgNO ₃ + KCI → AgCI + KNO ₃	Double	Replacement
10.	N ₂ + 3H ₂ → 2NH ₃	Synth	nesis
11.	2H ₂ O ₂ → 2H ₂ O + O ₂	Decomp	position
12.	$(NH_4)_2SO_4 + Ba(NO_3)_2 \rightarrow BaSO_4$	+ 2NH4NO3 DO	uble Replacement
13.	Mgl ₂ + Br ₂ → MgBr ₂ + l ₂	Single	Replacement

Synthesis

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RJ Alexander

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Record ,1901

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