

Worksheet #3: Naming Inorganic Compounds

1. Give the name for each of the following ionic compounds:

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- a. AlF_3 Aluminum Fluoride
- b. $Fe(OH)_2$ Iron(II) Hydroxide
- c. $Cu(NO_3)_2$ Copper(II) Nitrate
- d. $KMnO_4$ Potassium Permanganate
- e. Li_3PO_4 Lithium Phosphate
- f. HgS Mercury(II) Sulfide
- g. $Ca(CH_3COO)_2$ Calcium Acetate
- h. $Cr_2(CO_3)_3$ Chromium(III) Carbonate
- i. $K_2Cr_2O_7$ Potassium Dichromate
- j. $(NH_4)_2SO_4$ Ammonium Sulfate

Cu	(NO_3)
$+2$	-1
$\times 1$	$\times 2$
<hr/>	<hr/>
$+2$	-2

H	S
$+1$	-2
$\times 2$	$\times 1$
<hr/>	<hr/>
$+2$	-2

Cr	(CO_3)
$+3$	-2
$\times 2$	$\times 3$
<hr/>	<hr/>
$+6$	-6

These are all binary ionic compounds (+/-)

all of the underlined compounds are polyatomic ions → Memorize Them

2. Write the chemical formula for each of the following compounds:

- a. calcium oxide CaO
- b. potassium peroxide K_2O_2
- c. aluminum hydroxide $Al(OH)_3$
- d. iron(III) bromide $FeBr_3$

- e. mercury(II) bromide $HgBr_2$
 - f. iron(III) phosphate $Fe_3(PO_4)_2$
 - g. sodium hypobromite $NaBrO$
- ↑
uncommon polyatomic ion
"BrO⁻¹"

⑥ OH is a common cation +2

3. Give the name or chemical formula, as appropriate, for each of the following acids:

- a. HNO_3 Nitric Acid
- b. HBr Hydrobromic Acid
- c. H_3PO_4 Phosphoric Acid
phosphate $\frac{-3}{4}$ (the = oxy)

- d. hydrochloric acid $HClO$
 - e. iodic acid HIO_3
 - f. sulfurous acid H_2SO_3
- ↳ Break off of iodate as IO_3^-

↑ Bromine

4. Give the name or chemical formula, as appropriate, for each of the following molecular substances:

- a. SF_6 Sulfur hexafluoride
- b. IF_5 Iodine pentafluoride
- c. XeO_3 Xenon trioxide

- d. dinitrogen tetroxide N_2O_4
- e. hydrogen cyanide HCN
- f. tetraphosphorus hexasulfide P_4S_6

0 - Mono
1 - Di
2 - Di
3 - Tri
4 - Tetra
5 - Penta
6 - Hexa
7 - Hepta
8 - Octa
9 - Nona
10 - Deca

3

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Gen Chem Unit 3 Packet Answers

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C. Lum**

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