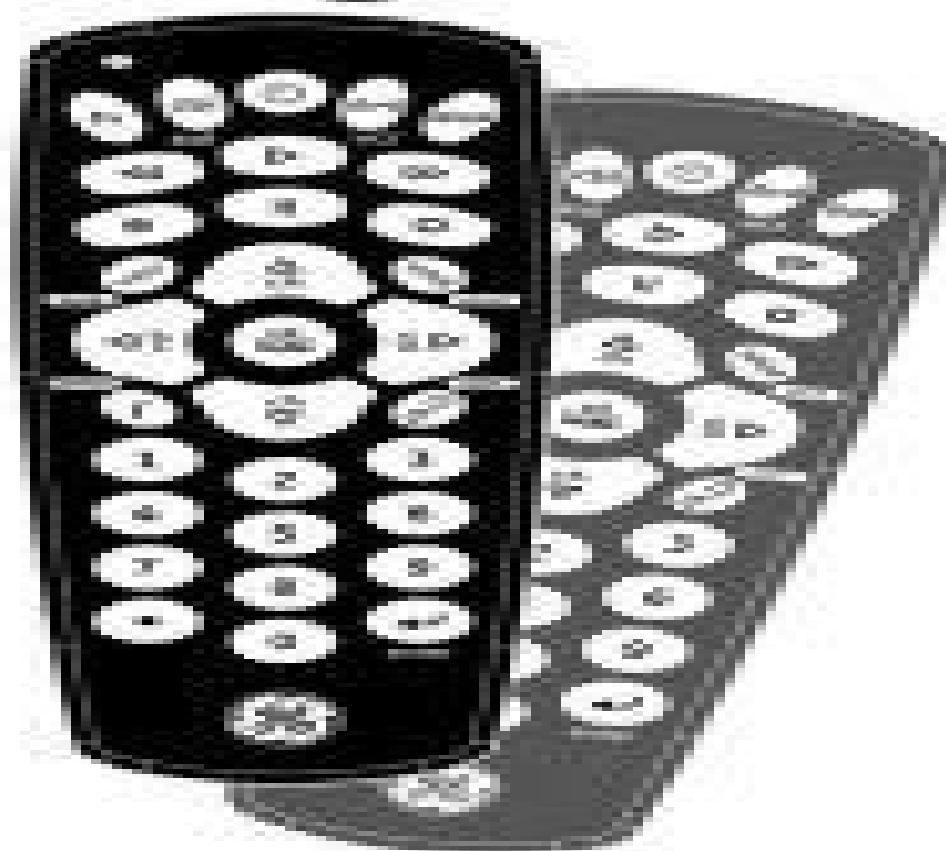
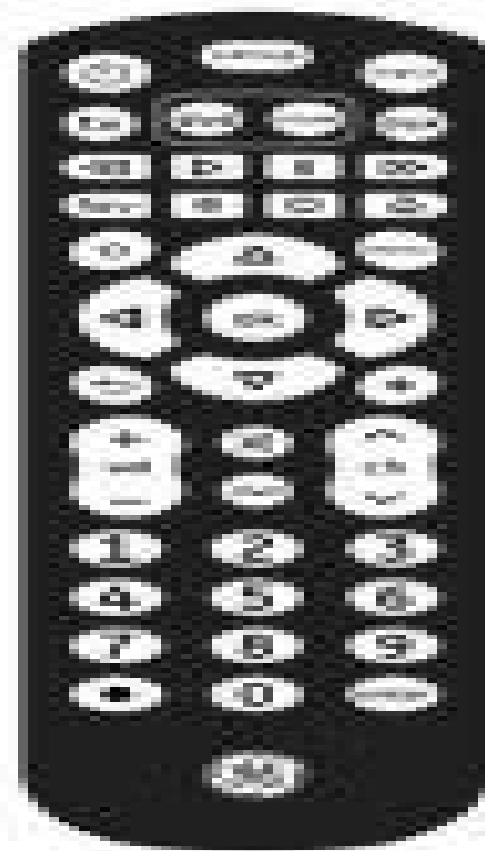


# GE UNIVERSAL Remote Controls Instruction Manuals Programming Guide



MODELS: 40081, 37527



MODELS: 33706, 33708,  
33707, 33709, 34456

# Ge Universal Remote Jc024 Instruction Manual

**G. E. Guibord, Canada. Indian and  
Northern Affairs Canada**



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**Instruction Manual for Use of VUCOM 3 Terminal** G. E. Guibord,Canada. Indian and Northern Affairs Canada,1977  
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web limiting reactant theoretical and percentage yields key 4  $\text{K}_2\text{S} + 2 \text{H}_2\text{O} + 4 \text{KOH} \rightarrow 3 \text{O}_2 + \text{g}$  if a reaction vessel contains 0 mol  $\text{K}_2\text{O}$  and 0 mol  $\text{H}_2\text{O}$  what is the limiting reactant?

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web limiting reactant reactant present in an amount lower than required by the reaction stoichiometry thus limiting the amount of product generated. Percent yield: measure of the efficiency of a reaction expressed as a percentage of the theoretical yield.

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  

$$8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8.3 \text{ mole } \text{H}_2\text{O}$$
  
 8.3 mole  $\text{H}_2\text{S} \times 18 \text{ g } \text{H}_2\text{O} / \text{mole } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$  is your theoretical yield so percent yield  

$$\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$$
  
 problems 6 & 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of  $\text{NH}_3$  can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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web limiting reagents problem one 1 for the balanced equation  

$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$
  
 if the reaction of 21.3 grams of  $\text{O}_2$  produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have  

$$21.3 \text{ g } \text{O}_2 \times \frac{1 \text{ mole } \text{O}_2}{32 \text{ g } \text{O}_2} = 0.666 \text{ moles } \text{O}_2$$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency mass of product obtained / calculated mass of the product expected  $\times 100$  in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction  

$$4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$$
  
 b determine the theoretical yield of  $\text{Al}_2\text{O}_3$   

$$160.0 \text{ g } \text{Al} \times \frac{1 \text{ mole } \text{Al}}{27 \text{ g } \text{Al}} \times \frac{2 \text{ mole } \text{Al}_2\text{O}_3}{4 \text{ mole } \text{Al}} \times 101.96 \text{ g } \text{Al}_2\text{O}_3 / \text{mole } \text{Al}_2\text{O}_3 = 302.3 \text{ g } \text{Al}_2\text{O}_3$$
  
 c determine the percent yield  

$$\frac{260.0 \text{ g } \text{Al}_2\text{O}_3}{302.3 \text{ g } \text{Al}_2\text{O}_3} \times 100 = 86.01\%$$

**7.2 limiting reagent and reaction yields chem 1114** - Jul 02 2022

web 7 Na<sub>2</sub>C<sub>2</sub>O<sub>4</sub> is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4.33 g of  $\text{H}_2$  are left over 13

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web the percent yield is calculated as follows text percent yield  $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  percent yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage text percent yield  $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  label 3 7 3

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