

Skills – Use $q = mc\Delta T$ to calculate the missing quantity, given the others.

Be sure to show all work required for the following calculations:

9. A 15.75g piece of iron absorbs 1006.75J of heat and its temperature changes from 25 to 175°C. Calculate the specific heat capacity of iron.

10. How many joules of heat are needed to raise the temp. of 10.0g of aluminum from 22 to 55°C, if the specific heat of aluminum is 0.90 J/g°C?

11. To what temperature will a 10.0g piece of glass raise if it absorbs 5275 J of heat and its specific heat capacity is 0.50 J/g°C? The initial temperature of the glass is 20°C.

12. Calculate the specific heat capacity of a piece of wood if 1500.0g of the wood absorbs 6.75x10⁵J, and its temp. changes from 32 to 57°C.

13. 100mL of 4.0°C water is heated until its temp. is 37°C. If the specific heat capacity of water is 4.18 J/g°C, calculate the amount of heat energy needed to cause this rise in temp. in joules AND calories.

14. 25.0g of an unknown substance is heated from 25 to 155°C, absorbing 455J in the process. Calculate the specific heat capacity of the unknown substance. Based on your answer determine the identity of the substance, using the chart on p329 of your text.

Skills – Calculate the enthalpy associated with a reaction.

15. Using the following balanced equation, calculate the thermal energy released (ΔH) when 45.0g of NH_4NO_3 dissolved in water?



16. Calculate the change in enthalpy (ΔH) when 1.00g of hydrogen peroxide decomposes.



Skills – Calculate the energy of a reaction using Hess's Law and identify if it is exothermic or endothermic.

Entropy Section Review Answers

Stephen R. Turns, Laura L. Pauley



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