

Atomic Structure – Study Guide

Grade 8 – Physical Science

Atom -- smallest particle of an element that has all the properties of that element.

3 main subatomic particles:

1. protons
2. neutrons
3. electrons



nucleus -- core of the atom, it is positively charged 99.9 % of the atoms mass is in the nucleus. Protons and neutrons are found in the **nucleus**.

1. Protons -- have a positive (+) charge. Found in the nucleus of the atom.

The **number of protons** determines what element an atom is.

Example: Carbon has 6 protons

Atomic Number = the **number of protons**, each element has it own unique number of protons.

2. Neutrons -- a neutral charge (no charge) found in the nucleus. These determine whether an element is radioactive.

Protons and neutrons are about *one atomic mass unit (amu)*. Electrons have a much smaller mass -- it takes almost 2000 electrons to equal 1 amu.

Atomic Mass = the total number of protons and neutrons.

Mass Number = Atomic Mass that is rounded.

To find just how many **neutrons** an atom has:

neutrons = atomic mass – atomic

Isotopes - atoms of the same element that have different number of neutrons.

Carbon-12

6 protons

6 neutrons

Carbon-14

6 protons

8 neutrons

NOTICE – the number of **protons DOES NOT** change if it is to be carbon -- only the number of neutrons.

3. Electrons - have a negative (-) charge. Found moving around the atom is a space called orbital's (shells).

Those electrons farthest away from the nucleus are called **valence electrons**, and are involved in the formation of chemical bonds.

In an **uncharged atom** the **number of protons is equal to the number of electrons**.

ion - when an atom gains or loses an *electron*. (has a charge)

If an atom gains electrons = negative (-) charge

If an atom loses electrons = positive (+) charge

How to find:

protons = atomic

neutrons = atomic mass – atomic

electrons = atomic # – charge (if any)

Matter And Atomic Structure Study Guide Answers

RS Peters



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Table of Contents Matter And Atomic Structure Study Guide Answers

1. Understanding the eBook Matter And Atomic Structure Study Guide Answers
 - The Rise of Digital Reading Matter And Atomic Structure Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Matter And Atomic Structure Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Matter And Atomic Structure Study Guide Answers
 - Personalized Recommendations
 - Matter And Atomic Structure Study Guide Answers User Reviews and Ratings
 - Matter And Atomic Structure Study Guide Answers and Bestseller Lists

5. Accessing Matter And Atomic Structure Study Guide Answers Free and Paid eBooks
 - Matter And Atomic Structure Study Guide Answers Public Domain eBooks
 - Matter And Atomic Structure Study Guide Answers eBook Subscription Services
 - Matter And Atomic Structure Study Guide Answers Budget-Friendly Options
6. Navigating Matter And Atomic Structure Study Guide Answers eBook Formats
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 - Matter And Atomic Structure Study Guide Answers Compatibility with Devices
 - Matter And Atomic Structure Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Matter And Atomic Structure Study Guide Answers
 - Highlighting and Note-Taking Matter And Atomic Structure Study Guide Answers
 - Interactive Elements Matter And Atomic Structure Study Guide Answers
8. Staying Engaged with Matter And Atomic Structure Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Matter And Atomic Structure Study Guide Answers
9. Balancing eBooks and Physical Books Matter And Atomic Structure Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Matter And Atomic Structure Study Guide Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Matter And Atomic Structure Study Guide Answers
 - Setting Reading Goals Matter And Atomic Structure Study Guide Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Matter And Atomic Structure Study Guide Answers
 - Fact-Checking eBook Content of Matter And Atomic Structure Study Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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