

Chapter 3: Matter and Energy – Study Guide

*Do not forget about the extra resources available to help you study and prepare for this test!

- Textbook pages 81- 93 (answers to odd questions in the back of the book)
- Mastering Chemistry Study Area: <http://www.pearsonmylabandmastering.com/northamerica/>
- Miss Marnik's Website: <http://www.northallegheny.org/Page/16665>
- Labs, problem packets, and homework assignments

Section 1 – Classifying Matter: (text pages 57-62)

Need to be able to:

- Define matter, atoms, molecules, element, compound, and mixture.
- Classify matter as and distinguish between solid, liquid, or gas.
- Classify matter as element, compound, or mixture.

Practice Problems:

1. Match the following terms with the phrases below:

- | | |
|--|--------------------------|
| a. State of matter in which atoms or molecules pack close to each other in fixed locations | 1. heterogeneous mixture |
| b. State of matter in which atoms or molecules are packed close together but are free to move around and by each other | 2. homogeneous mixture |
| c. Substance composed of only one type of atom or molecule | 3. liquid |
| d. Matter composed of two or more different types of atoms or molecules combined in variable proportions | 4. mixture |
| e. Mixture that has two or more regions with different compositions | 5. pure substance |
| f. Mixture that has the same composition throughout | 6. solid |

2. Which of the following can be classified as matter?

- | | | | |
|----------|------------|---------------------|-----------|
| a. atoms | c. wind | e. potential energy | g. carbon |
| b. gas | d. liquids | f. temperature | h. light |

3. Classify each type of matter as a pure substance or a mixture. If it is a pure substance, classify it as an element or a compound. If it is a mixture, classify it as homogeneous or heterogeneous.

- | | |
|----------------------------------|---------------|
| a. limestone (CaCO_3) | e. Chex mix |
| b. orange juice with pulp | f. Table salt |
| c. helium | g. Brass |
| d. Pacific ocean | h. oxygen |

Matter And Energy Study Guide

Tobias Bleicker



Matter And Energy Study Guide:

Matter and Energy ,2005 **Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual** Leo J.

Malone,Theodore O. Dolter,2012-01-03 The 9th edition of Malone s Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter Matter and

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carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH₄ from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

Chemical Principles Study Guide/Solutions Manual John Krenos, Joseph Potenza, Carl Hoeger, 2007-01-18 Written for general chemistry courses Chemical Principles helps students develop chemical insight by showing the connection between chemical principles and their applications

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