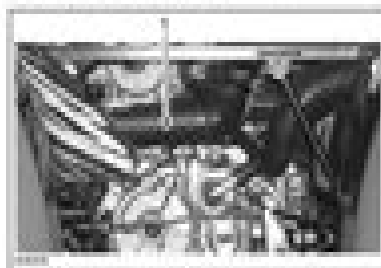


- rear exhaust shield no. 4 (this will provide water access to exhaust spring)



1. Rear exhaust shield
- seat support bracket
 - exhaust spring no. 5
 - exhaust pipe no. 6

Disconnect

- wires to starter
- water temperature sensor
- oil pressure sensor
- lower coolant hose from engine
- large CVT breather hose



1. Water temperature sensor
2. Oil pressure sensor
3. Lower coolant hose
4. CVT breather hose

On left side of vehicle, remove the following:

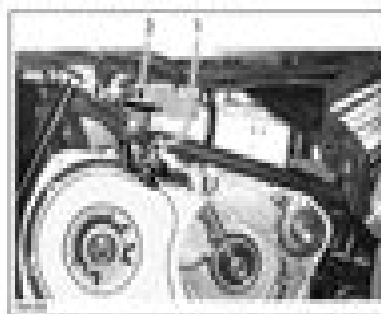
- LH engine side panel (refer to BODY)
- CVT breather box



1. Starter box
- footrest (refer to BODY)
 - CVT cover (refer to CVT)
 - drive and driven pulleys (refer to CVT)

Disconnect hose from fuel valve to fuel pump

Remove fuel valve bracket from frame and place valve and hoses off to one side



1. Fuel valve
2. Fuel valve bracket

Disconnect magnets, gear box and speed sensor connectors.



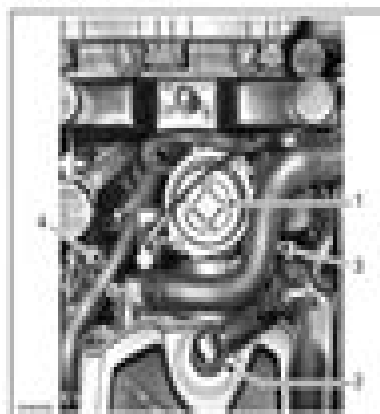
1. Magnet connector
2. Gear box connector
3. Speed sensor connector

From the top side, do the following:

Remove carburetor (refer to CARBURETOR AND FUEL PUMP)

Disconnect

- the ignition coil and remove it
- upper coolant hose



1. Carburetor
2. Ignition coil
3. Upper coolant hose
4. Breather hose

Remove cylinder head breather hose. Refer to CYLINDER AND HEAD

Remove upper engine mount assembly no. 7 and no. 8



Insert engine lifting tool onto exhaust mounting studs.



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Rachel S Tattersall



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the boiling point of an organic liquid is the temperature at which its vapour pressure equals the atmospheric pressure over the liquid or it is the temperature at which the vapour and liquid phases are in equilibrium at a given pressure the boiling point is considered as a criterion of purity of a compound and is useful for identification of

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given that the boiling points of ethanol and methanol 78 °C and 64 °C respectively are far off in a significant way this seems like a valid conclusion conclusion thermometer calibrations are necessary in order to collect reliable information as precise calculations are imperative to gathering correct results

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melting point 1 a push the open end of a capillary tube into the powdered acetamide b move the powder to the closed end of the capillary tube by tapping it on the table repeat until the powdered acetamide occupies 1-2 mm of the capillary tube end c

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discussion and conclusion the technique recrystallization was used because it is a simple method to purify a solid compound the percentage yield for benzoic acid and acetanilide acid being high indicated the initial sample had a low concentration of impurities with the melting point being in a specific

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introduction in this lab we will be using boiling points melting points and infrared spectrometry to identify unknown

compounds boiling points are affected by three key factors the first being intermolecular forces inter molecular affect boiling point based on bond strength

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determination of boiling points introduction the boiling point of a compound is the temperature at which it changes from a liquid to a gas this is a physical property often used to identify substances or to check the purity of the compound it

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conclusion after completing this experiment the melting point of three different unknown compounds was obtained using the mel temp apparatus compound a had a melting point range of 114.6-115.5 c thus revealing it to be acetanilide compound b had a melting point range of 122.6-124.5 c thus

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melting point can be used as the identity and purity of an organic solid however it is not the same as the boiling point the boiling point is a characteristic unique of pure liquids the boiling point is the temperature at which the vapor pressure of a liquid is equal to the atmospheric pressure

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50 c so in order to convert an experimental measurement taken in calgary which is at higher altitude than sea level and so is at a lower pressure to that reported for sea level higher pressure one needs to add a correction factor since the boiling point at sea level is higher than that at higher altitudes

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