

Stoichiometry

Answer Key

Vocabulary: Avogadro's number, balanced equation, cancel, coefficient, dimensional analysis, molar mass, mole, molecular mass, stoichiometry

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

[Note: The purpose of these questions is to activate prior knowledge and get students thinking. Students are not expected to know the answers to the Prior Knowledge Questions.]

1. A 250 mL glass of orange juice contains 22 grams of sugar. How much sugar is in a two-liter (2,000 mL) bottle of orange juice? *176 grams of sugar*
2. It requires two sticks of butter to make a batch of 20 cookies. How much butter will it take to make 150 cookies? *15 sticks of butter*

Gizmo Warm-up

Just as a cook follows a recipe to decide how much of each ingredient to add, a chemist uses **stoichiometry** to determine the amounts of substances involved in chemical reactions. The Stoichiometry Gizmo™ allows you to try your hand at figuring out the amounts of reactants and products that take part in a chemical reaction.

To begin, check that this equation is shown:



1. Look at the **coefficients** (such as the "3" in 3CO) in front of each substance in the equation. The coefficients tell you how many molecules or atoms take part in a chemical reaction. In the spaces below, list the number of each molecule or atom in the equation:

Fe_2O_3 *1*

CO *3*

Fe *2*

CO_2 *3*

2. In a **balanced equation**, the same number of each kind of atom is shown on each side of the equation. Calculate the number of iron (Fe), oxygen (O), and carbon atoms (C).

Reactants

Iron: *2*

Oxygen: *6*

Carbon: *3*

Products

Iron: *2*

Oxygen: *6*

Carbon: *3*

$$\text{Fe}_2\text{O}_3 (\text{s}) + 3\text{CO} (\text{g}) \rightarrow 2\text{Fe} (\text{s}) + 3\text{CO}_2 (\text{g})$$

How many moles of carbon monoxide (CO) are required to react completely with 1.75 moles of iron (III)-oxide (Fe_2O_3)?

Problems solved: 0

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Kaplan Test Prep



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Table of Contents Lesson 28 Stoichiometry Answer Key

1. Understanding the eBook Lesson 28 Stoichiometry Answer Key
 - The Rise of Digital Reading Lesson 28 Stoichiometry Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Lesson 28 Stoichiometry Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Lesson 28 Stoichiometry Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Lesson 28 Stoichiometry Answer Key
 - Personalized Recommendations
 - Lesson 28 Stoichiometry Answer Key User Reviews and Ratings
 - Lesson 28 Stoichiometry Answer Key and Bestseller Lists
5. Accessing Lesson 28 Stoichiometry Answer Key Free and Paid eBooks
 - Lesson 28 Stoichiometry Answer Key Public Domain eBooks
 - Lesson 28 Stoichiometry Answer Key eBook Subscription Services

- Lesson 28 Stoichiometry Answer Key Budget-Friendly Options
- 6. Navigating Lesson 28 Stoichiometry Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Lesson 28 Stoichiometry Answer Key Compatibility with Devices
 - Lesson 28 Stoichiometry Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Lesson 28 Stoichiometry Answer Key
 - Highlighting and Note-Taking Lesson 28 Stoichiometry Answer Key
 - Interactive Elements Lesson 28 Stoichiometry Answer Key
- 8. Staying Engaged with Lesson 28 Stoichiometry Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Lesson 28 Stoichiometry Answer Key
- 9. Balancing eBooks and Physical Books Lesson 28 Stoichiometry Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Lesson 28 Stoichiometry Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Lesson 28 Stoichiometry Answer Key
 - Setting Reading Goals Lesson 28 Stoichiometry Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Lesson 28 Stoichiometry Answer Key
 - Fact-Checking eBook Content of Lesson 28 Stoichiometry Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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