

Student Exploration: Covalent Bonds

Vocabulary: covalent bond, diatomic molecule, Lewis diagram, molecule, noble gases, nonmetal, octet rule, shell, valence, valence electron

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

1. There are eight markers in a full set, but Flora and Frank each only have seven markers. Flora is missing the red marker, and Frank is missing the blue marker.

What can they do so that each has a full set of markers? *Frank can share his red marker with Flora and Flora can share her blue marker with Frank.*
2. Otto and Olivia each have six markers. Otto is missing the purple and green markers, and Olivia is missing the black and brown markers. What can they do so that each has a full set? *Olivia can share her purple and green markers with Otto and Otto can share his black and brown markers with Olivia.*

Gizmo Warm-up

Just like the students described above, **nonmetal** atoms can share electrons. As you will see in the *Covalent Bonds* Gizmo, atoms form bonds in this way.



To begin, check that **Fluorine** is selected from the **Select a substance** menu. Click **Play** () to see the electrons orbiting the nucleus of each atom.

1. The outermost electrons in each atom are called **valence electrons**. How many valence electrons does each fluorine atom have? *Each fluorine atom has seven valence electrons.*
2. Click **Pause** (). Drag an electron from the left atom to the right atom. Click **Play**. What happens? *The electrons that was dragged now orbits both atoms.*
3. Click **Pause**, drag an electron from the right atom to the left, and then click **Play**. What happens now? *A pair of electrons now orbit both atoms.*

Lesson 20 Covalent Bonds Answer Key

SJ Ball



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Table of Contents Lesson 20 Covalent Bonds Answer Key

1. Understanding the eBook Lesson 20 Covalent Bonds Answer Key
 - The Rise of Digital Reading Lesson 20 Covalent Bonds Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Lesson 20 Covalent Bonds Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Lesson 20 Covalent Bonds Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Lesson 20 Covalent Bonds Answer Key
 - Personalized Recommendations
 - Lesson 20 Covalent Bonds Answer Key User Reviews and Ratings
 - Lesson 20 Covalent Bonds Answer Key and Bestseller Lists
5. Accessing Lesson 20 Covalent Bonds Answer Key Free and Paid eBooks
 - Lesson 20 Covalent Bonds Answer Key Public Domain eBooks
 - Lesson 20 Covalent Bonds Answer Key eBook Subscription Services

- Lesson 20 Covalent Bonds Answer Key Budget-Friendly Options
- 6. Navigating Lesson 20 Covalent Bonds Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Lesson 20 Covalent Bonds Answer Key Compatibility with Devices
 - Lesson 20 Covalent Bonds Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Lesson 20 Covalent Bonds Answer Key
 - Highlighting and Note-Taking Lesson 20 Covalent Bonds Answer Key
 - Interactive Elements Lesson 20 Covalent Bonds Answer Key
- 8. Staying Engaged with Lesson 20 Covalent Bonds Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Lesson 20 Covalent Bonds Answer Key
- 9. Balancing eBooks and Physical Books Lesson 20 Covalent Bonds Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Lesson 20 Covalent Bonds Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Lesson 20 Covalent Bonds Answer Key
 - Setting Reading Goals Lesson 20 Covalent Bonds Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Lesson 20 Covalent Bonds Answer Key
 - Fact-Checking eBook Content of Lesson 20 Covalent Bonds Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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