

Chemical Bonds

Strand	Matter	
Topic	Investigating ionic and covalent bonds	
Primary SOL	PS.4	The student will investigate and understand the organization and use of the periodic table of elements to obtain information. Key concepts include c) formation of compounds through ionic and covalent bonding.
Related SOL	PS.1	The student will demonstrate an understanding of scientific reasoning, logic, and the nature of science by planning and conducting investigations in which k) research methods are used to investigate practical problems and questions; m) models and simulations are constructed and used to illustrate and explain phenomena
	PS.2	The student will investigate and understand the nature of matter. Key concepts include b) elements, compounds, mixtures, acids, bases, and salts

Background Information

Compounds are formed when two or more elements chemically combine. Elements form compounds by filling their outermost energy level of electrons. Once an atom's outermost energy level is full, the atom becomes relatively stable. There are two types of compounds covered in this lesson—ionic and covalent.

One way an element can fill its outermost energy level is by either gaining or losing valence electrons. Two or more ions held next to each other by electrical attraction creates an ionic compound. One of the ions in an ionic compound has a positive charge (called a cation) and the other has a negative charge (anion). Cations are usually metal ions and anions are either nonmetals or polyatomic ions (ions with more than one atom).

Atoms can fill their outermost energy level by sharing electrons creating a covalent bond. Instead of metals bonding with nonmetals as in ionic compounds, nonmetals, and metalloids covalently bond only with other nonmetals and metalloids. Molecules are compounds where the atoms are bonded covalently.

Materials

- Resource materials on ionic and covalent bonding
- Construction or white paper
- Drawing utensils

Vocabulary

compound, covalent bond, ion, ionic bond

Lesson 12 Chemical Bonds

Terry C. Jones



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Table of Contents Lesson 12 Chemical Bonds

1. Understanding the eBook Lesson 12 Chemical Bonds
 - The Rise of Digital Reading Lesson 12 Chemical Bonds
 - Advantages of eBooks Over Traditional Books
2. Identifying Lesson 12 Chemical Bonds
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Lesson 12 Chemical Bonds
 - User-Friendly Interface
4. Exploring eBook Recommendations from Lesson 12 Chemical Bonds
 - Personalized Recommendations

- Lesson 12 Chemical Bonds User Reviews and Ratings
- Lesson 12 Chemical Bonds and Bestseller Lists
- 5. Accessing Lesson 12 Chemical Bonds Free and Paid eBooks
 - Lesson 12 Chemical Bonds Public Domain eBooks
 - Lesson 12 Chemical Bonds eBook Subscription Services
 - Lesson 12 Chemical Bonds Budget-Friendly Options
- 6. Navigating Lesson 12 Chemical Bonds eBook Formats
 - ePub, PDF, MOBI, and More
 - Lesson 12 Chemical Bonds Compatibility with Devices
 - Lesson 12 Chemical Bonds Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Lesson 12 Chemical Bonds
 - Highlighting and Note-Taking Lesson 12 Chemical Bonds
 - Interactive Elements Lesson 12 Chemical Bonds
- 8. Staying Engaged with Lesson 12 Chemical Bonds
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Lesson 12 Chemical Bonds
- 9. Balancing eBooks and Physical Books Lesson 12 Chemical Bonds
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Lesson 12 Chemical Bonds
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Lesson 12 Chemical Bonds
 - Setting Reading Goals Lesson 12 Chemical Bonds
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Lesson 12 Chemical Bonds
 - Fact-Checking eBook Content of Lesson 12 Chemical Bonds

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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