

7 Ionic and Metallic Bonding

Big idea BONDING AND INTERACTIONS

7.1 Ions

For students using the Foundation edition, assign problems 1, 3–5, 7–12, 14, 15, 18–20.

Essential Understanding Ions form when atoms gain or lose valence electrons, becoming electrically charged.

Lesson Summary

Valence Electrons Valence electrons are the electrons in the outermost occupied energy level and are involved in ion formation.

- ▶ For a representative element, the group number equals the number of valence electrons the atom contains.
- ▶ An electron dot structure shows the symbol of the element and its valence electrons.
- ▶ Atoms tend to gain or lose the number of electrons that will provide the atom with a noble gas electron configuration.

Formation of Cations Cations are positively charged ions formed when an atom loses one or more valence electrons.

- ▶ Atoms and the cations formed from them have different properties.
- ▶ Elements in Group 1 form cations with a charge of $1+$, and those in Group 2 form cations with a charge of $2+$.
- ▶ Many transition metals form more than one cation and do not follow the octet rule.

Formation of Anions Anions are negatively charged ions formed when an atom gains one or more valence electrons.

- ▶ Commonly, the name of an anion ends in *-ide*.
- ▶ Anions form from nonmetallic elements.
- ▶ The anions formed from halogens are known as halides.

After reading Lesson 7.1, answer the following questions.

Valence Electrons

1. What are valence electrons?

Valence electrons are the electrons in the highest occupied energy level of an element's atoms.

2. The valence electrons largely determine the *chemical properties* of an element and are usually the only electrons used in *chemical bonds*.

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Whispering the Techniques of Language: An Psychological Quest through **Lesson 6 1 Reading And Study Workbook Chemistry**

In a digitally-driven world wherever monitors reign great and instant communication drowns out the subtleties of language, the profound strategies and psychological nuances concealed within phrases often get unheard. However, nestled within the pages of **Lesson 6 1 Reading And Study Workbook Chemistry** a interesting fictional treasure pulsating with organic emotions, lies an extraordinary journey waiting to be undertaken. Written by a talented wordsmith, that enchanting opus invites viewers on an introspective trip, gently unraveling the veiled truths and profound influence resonating within the very fabric of each and every word. Within the mental depths with this moving review, we shall embark upon a genuine exploration of the book is primary themes, dissect its captivating publishing fashion, and fail to the effective resonance it evokes heavy within the recesses of readers hearts.

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