

The frequency of the photon at the convergence limit in the highest energy emission series of hydrogen is 3.28×10^{15} Hz.

Calculate the first ionization energy of hydrogen, giving your answer in kJ mol^{-1} .



Turn around

288 | ? HL Topic 12.1 | Electrons in atoms

Step 1: Use the following formula:

$$E = h\nu$$

Step 2: Multiply Planck's constant by the frequency:

$$6.63 \times 10^{-34} \times 3.28 \times 10^{15} = 2.17 \times 10^{-18} \text{ J}$$

Step 3: Multiply the energy by Avogadro's number:

$$2.17 \times 10^{-18} \times 6.02 \times 10^{23} = 1,310,000 \text{ J mol}^{-1}$$

Step 4: Divide by 1000 to convert to kJ:

$$\frac{1,310,000 \text{ J mol}^{-1}}{1,000} = 1,310 \text{ kJ mol}^{-1}$$

Planck's constant (h) is given in the data booklet as is the equation $E = h\nu$. Don't forget to convert to the units asked for in the question.

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Unveiling the Magic of Words: A Report on "**Ib Hl Chemistry Data Let 2014**"

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