

Element	Charge	Radius / pm	Melting point (K)
Li	+1	152	454
Na	+1	186	371
K	+1	231	337
Rb	+1	244	312
Cs	+1	262	302

Table 4.4: Melting Points of the Alkali Metals

You can get two points for one comment – “charge density”. As the ion gets larger, the charge is diluted over the increased surface of the ionic sphere. As the charge density decreases, the force of attraction between the ion and the delocalized electrons is lowered, therefore the melting point lowers.

## Covalent Bonding

Covalent bonding is the attractive force between two nuclei and shared pairs of electrons, due to overlap of their orbitals. It occurs when there is a high average electronegativity and a low difference in electronegativity between two atoms.

If one pair is shared, it is a single bond. Double bonds occur when two pairs of electrons are shared; triple is three pairs.

Like all other types of bonding, the size of the atom and charge of the nucleus and the number of electrons involved in the bond will be important factors.

Covalent bonding arises when there is a low difference of electronegativity, but a high average electronegativity – see the bonding triangle in your data book.

## Covalent Formulae

The following molecules are common gases in chemistry questions.

Name	Formula	Lewis Structure	comment
oxygen	O <sub>2</sub>		~20% of air
nitrogen	N <sub>2</sub>		~80% of air
carbon dioxide	CO <sub>2</sub>		you breathe out, plants absorb
hydrogen cyanide	HCN		very poisonous
ethane	C <sub>2</sub> H <sub>6</sub>		a flammable hydrocarbon gas
ethene	C <sub>2</sub> H <sub>4</sub>		contains a double bond that is reactive
ethyne	C <sub>2</sub> H <sub>2</sub>		contains a triple bond

Table 4.5: Some common covalent molecules

However, we need a way to draw structures for any molecule given, not just memorizing the common ones.

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