

⇒ Heisenberg uncertainty principle :

$$\Delta x \cdot \Delta p \geq \frac{h}{4\pi}$$

⇒ Number of orbitals in a shell =  $n^2$

⇒ Number of electrons in a shell =  $2n^2$

⇒ Number of radial nodes =  $(n - l - 1)$

### Chemical Bonding

$$H = \frac{1}{2}[V + X - C + A]$$

(H = number of Hybrid orbitals)

$$\Rightarrow \text{Bond order} = \frac{1}{2}[N_b - N_a]$$

### States of matter

⇒ Boyle's law :  $P_1 V_1 = P_2 V_2$

⇒ Charles law :  $\frac{V_1}{T_1} = \frac{V_2}{T_2}$

⇒ Ideal gas equation :  $PV = nRT$

Dalton's law :  $\Rightarrow P_{\text{total}} = P_1 + P_2 + P_3 + \dots$

$$\Rightarrow d_{\text{gas}} = \frac{PM}{RT}$$

$$\text{Graham's Law} \Rightarrow \frac{r_1}{r_2} = \sqrt{\frac{M_2}{M_1}}$$

⇒ Vander waal's equation :  $\left(P + \frac{an^2}{V^2}\right)(V - nb) = nRT$

⇒  $C_p - C_v = nRT$

⇒ Kinetic gas equation :  $PV = \frac{1}{3}Mu^2$

### Chemical Thermodynamics

⇒  $\Delta U = q_p$ ,  $\Delta H = q_p$

⇒  $\Delta H = \Delta U + \Delta n_g RT$

⇒  $C = \frac{q}{\Delta T}$  (C = Heat capacity)

⇒  $\Delta U = q + w$  (First law)

⇒  $\Delta H = \sum \Delta_f H^\circ_{\text{products}} - \sum \Delta_f H^\circ_{\text{reactants}}$

⇒  $\Delta G = \Delta H - T\Delta S$

⇒  $\Delta G^\circ = -RT \ln K$

### Equilibrium

$$\Rightarrow K_{eq} = \frac{K_c}{K_p}$$

⇒  $K_p = K_c(RT)^{\Delta n}$

⇒  $K_w = [H^+][OH^-]$

⇒  $pH = -\log [H^+]$

⇒  $pK_w = pH + pOH$

⇒  $K_a = C\alpha^2$

⇒  $K_w = K_a \times K_b$

⇒  $K_w = [A^-]_e [B^+]$

### Solid State

$$\Rightarrow \text{Density} = \frac{Z \times M}{a^3 \times N_A} \text{ g/cm}^3$$

⇒ Packing fraction : Simple cubic =  $\frac{\pi}{6}$ , FCC =  $\frac{\sqrt{2}\pi}{6}$ , BCC =  $\frac{\sqrt{3}\pi}{8}$

⇒ Bragg's equation  $d = n\lambda \sin \theta$

### Solution

⇒ Henry's law  $X_{\text{gas}} = K_H P$

⇒ Raoult's law : v.p. of solution =  $x_A P_A^\circ + x_B P_B^\circ$

⇒ Relative lowering of V.P. :  $\frac{P_A^\circ - P_A}{P_A^\circ} = X_{\text{solvent}}$

⇒  $\Delta T_b = i K_b m$

⇒  $\Delta T_f = i K_f m$

⇒  $\pi = CRT$

⇒  $i = \frac{\text{no. of particle of solute in solution}}{\text{no. of particles of pure solute}}$

### Electrochemistry

⇒  $E^\circ_{\text{cell}} = E^\circ_{\text{cathode}} - E^\circ_{\text{anode}}$

⇒  $\Delta G^\circ = -nF E^\circ_{\text{cell}}$

# Hsc Chemistry Formula Sheet

**C. M. Roebuck**



## Hsc Chemistry Formula Sheet:

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