

**SECTION - I**

**Q. 1. Select and write the most appropriate answers from the given alternatives for each sub-question :**

[7]

- (i) The Arrhenius equation is \_\_\_\_\_.
- (a)  $k = A e^{\frac{RT}{E_a}}$  (b)  $A = k e^{\frac{-E_a}{RT}}$
- (c)  $k = A e^{\frac{-RT}{E_a}}$  (d)  $k = A e^{\frac{-K_a}{RT}}$
- (ii) If the enthalpy of vaporisation of water at 100°C is 186.5 J . mol<sup>-1</sup>, the entropy of vaporization will be \_\_\_\_\_.
- (a) 4.0 J . K<sup>-1</sup>. mol<sup>-1</sup> (b) 3.0 J . K<sup>-1</sup>. mol<sup>-1</sup>
- (c) 1.5 J . K<sup>-1</sup>. mol<sup>-1</sup> (d) 0.5 J . K<sup>-1</sup>. mol<sup>-1</sup>
- (iii) The atomicity of sulphur in orthorhombic sulphur is \_\_\_\_\_.
- (a) 8 (b) 6 (c) 4 (d) 2
- (iv) The major binding force in diamond is \_\_\_\_\_.
- (a) Covalent bond (b) Ionic bond
- (c) Metallic bond (d) Co-ordinate covalent bond
- (v) The boiling point of water at high altitude is low, because \_\_\_\_\_.
- (a) the temperature is low. (b) the atmospheric pressure is low.
- (c) the temperature is high. (d) the atmospheric pressure is high.
- (vi) The molar conductivity of cation and anion of salt BA are 180 and 220 mhos respectively. The molar conductivity of salt BA at infinite dilution is \_\_\_\_\_.
- (a) 90 mhos.cm<sup>2</sup>. (b) 110 mhos.cm<sup>2</sup>.mol<sup>-1</sup>
- (c) 200 mhos.cm<sup>2</sup>.mol<sup>-1</sup> (d) 400 mhos.cm<sup>2</sup>.mol<sup>-1</sup>
- (vii) What is the process in which concentrated ore is reduced to the corresponding metal by heating at high temperature with a reducing agent?
- (a) Polling (b) Pyrometallurgy (c) Hydrometallurgy (d) Calcination

**Q. 2. Answer any THREE of the following :**

[9]

- (i) Describe anomalous behaviour of oxygen as compared with other elements of group 16 with reference to –
- (a) Magnetic property (b) Oxidation state (c) Hydrides
- (ii) What is the value of  $\Delta S_{\text{surr.}}$  for the following reaction at 298 K –
- $$6\text{CO}_{2(g)} + 6\text{H}_2\text{O}_{(l)} \longrightarrow \text{C}_6\text{H}_{12}\text{O}_{6(s)} + 6\text{O}_{2(g)}$$
- Given that :  $\Delta G^\circ = 2879 \text{ kJ mol}^{-1}$ ,  $\Delta S = -210 \text{ JK}^{-1} \text{ mol}^{-1}$ .
- (iii) Sucrose decomposes in acid solution to give glucose and fructose according to the first order rate law. The half life of the reaction is 3 hours. Calculate fraction of sucrose which will remain after 8 hours.
- (iv) A solution containing 0.73 g of camphor (molar mass 152 g . mol<sup>-1</sup>) in 36.8 g of acetone (boiling point 56.3°C) boils at 56.55°C. A solution of 0.564 g of unknown compound in the same weight of acetone boils at 56.46°C. Calculate the molar mass of the unknown compound.

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**JL Elias**



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### **Table of Contents Hsc 2014 Chemistry 2nd Paper Question**

1. Understanding the eBook Hsc 2014 Chemistry 2nd Paper Question
  - The Rise of Digital Reading Hsc 2014 Chemistry 2nd Paper Question
  - Advantages of eBooks Over Traditional Books
2. Identifying Hsc 2014 Chemistry 2nd Paper Question
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Hsc 2014 Chemistry 2nd Paper Question
  - User-Friendly Interface
4. Exploring eBook Recommendations from Hsc 2014 Chemistry 2nd Paper Question

- Personalized Recommendations
- Hsc 2014 Chemistry 2nd Paper Question User Reviews and Ratings
- Hsc 2014 Chemistry 2nd Paper Question and Bestseller Lists
- 5. Accessing Hsc 2014 Chemistry 2nd Paper Question Free and Paid eBooks
  - Hsc 2014 Chemistry 2nd Paper Question Public Domain eBooks
  - Hsc 2014 Chemistry 2nd Paper Question eBook Subscription Services
  - Hsc 2014 Chemistry 2nd Paper Question Budget-Friendly Options
- 6. Navigating Hsc 2014 Chemistry 2nd Paper Question eBook Formats
  - ePub, PDF, MOBI, and More
  - Hsc 2014 Chemistry 2nd Paper Question Compatibility with Devices
  - Hsc 2014 Chemistry 2nd Paper Question Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Hsc 2014 Chemistry 2nd Paper Question
  - Highlighting and Note-Taking Hsc 2014 Chemistry 2nd Paper Question
  - Interactive Elements Hsc 2014 Chemistry 2nd Paper Question
- 8. Staying Engaged with Hsc 2014 Chemistry 2nd Paper Question
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Hsc 2014 Chemistry 2nd Paper Question
- 9. Balancing eBooks and Physical Books Hsc 2014 Chemistry 2nd Paper Question
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Hsc 2014 Chemistry 2nd Paper Question
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Hsc 2014 Chemistry 2nd Paper Question
  - Setting Reading Goals Hsc 2014 Chemistry 2nd Paper Question
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Hsc 2014 Chemistry 2nd Paper Question

- Fact-Checking eBook Content of Hsc 2014 Chemistry 2nd Paper Question
- Distinguishing Credible Sources

### 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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