

Basic Stoichiometry

Why?

In this activity we will address the question:

How do I convert between moles of different chemical species in a given reaction?

Model 1



2 mol A produces 5 mol C

3 mol B produces 4 mol D

2 mol A reacts with 3 mol B

4 mol A produces 10 mol C

6 mol B produces 8 mol D

6 mol A reacts with 9 mol B

- What do the coefficients of a balanced equation represent?
- If 6 mol A are reacted, how many mol D are formed? Explain IN WORDS how you figured this out.
- If 15 mol B are reacted, how many mol C are formed? Show your work.
- If 8 mol of A are reacted, how many mol B are used? Show your work.

Model 2



If 3 mol A react, how many mol C are produced?

$$3 \text{ mol A} \times \frac{5 \text{ mol C}}{2 \text{ mol A}} = 7.5 \text{ mol C}$$

$$\text{In general: } \text{mol known} \times \frac{\text{mol unknown}}{\text{mol known}} = \text{mol unknown}$$

- If 21.0 mol A are reacted, how many mol C are produced? Using dimensional analysis, show how you calculated your answer.
- If 21.0 mol B are reacted, how many mol A must have reacted?

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