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web consider the thermochemical equation below $4 \text{ NO(g)} + 6 \text{ H}_2\text{O(l)} + 4 \text{ NH}_3\text{(g)} + 5 \text{ O}_2\text{(g)} \rightarrow 1170 \text{ kJ}$ a is this reaction endothermic or exothermic endothermic b how much heat will be absorbed or released if 10 g of no is reacted with 10 g h₂ o

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web heats of fusion and vaporization determines what determines how much energy is required to change a substance s phase what is the heat of fusion formula $h_n = h_{\text{fus}}$ what is the heat of vaporization formula $h_n = h_{\text{vap}}$ what is on the x axis of the heating curve time

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web sep 5 2023 3 carry out a thermochemical calculation on all these species at their optimized geometries 4 sum up the enthalpies free energy of the products and the reagents keeping into account the

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web answer a 1 explanation the entire reaction was multiplied by 2 therefore Δh must also be halved 3 answer c explanation option c is incorrect because 566 kJ mol is the Δh per mole of o₂ and not for 2 moles even though options b and d are

expressed in grams they are still correct because 56

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web thermochemistry practice worksheet answer key what is the heat change when 472 g of carbon reacts with excess O_2 according to the following equation how much heat energy must be expelled by the body to rid the body of one pound of fat
 $2S + 3O_2 \rightarrow 2SO_3$ $\Delta H = -7914 \text{ kJ}$ is this endothermic or exothermic?

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web $q = 140 \text{ g} \times 4.184 \text{ J/g} \times 16.0^\circ\text{C}$ $q = 9372 \text{ J}$ $35.5 \text{ g} \times 1 \text{ mole KCN} \times 0.5453 \text{ moles}$ $65.1 \text{ g} \times 0.5453 \text{ mole moles}$ 5 when a 15 g sample of $SnCl_2$ dissolves in 80 g of water in a calorimeter the temperature drops from 42°C to 15°C calculate the ΔH for the process

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