

II PUC CHEMISTRY SYLLABUS BLOW-UP

UNIT-I

Solid State

8 hrs

General characteristics of solids: amorphous and crystalline solids – examples, differences. Classification of crystalline solids – based on binding forces: molecular solids – (non-polar, polar, H-bonded), ionic solids, metallic solids, covalent or network solids – examples for all.

Definitions – lattice point, crystal lattice, unit cell, coordination number. Parameters of a unit cell, names of seven crystal systems, calculation of number of atoms in a cubic unit cell – simple cubic, fcc, bcc. Close packing in two-dimensional and three dimensional lattices – brief information, voids – types of voids, tetrahedral and octahedral and their relative numbers, calculation of the formula of the compounds based on the number of voids filled. Packing in solids – calculation of packing efficiency- fcc/bcc, bcc, simple cubic. Formula to calculate the density of the unit cell to be assumed- use the formula to calculate a , d , z , M , N_A . Numerical problems.

Point defects-types, a brief account of Frenkel and Schottky defects, metal excess defect and metal deficiency defect with examples.

Electrical properties: classification into conductors, insulators and semiconductors – their comparison based on band theory of metals, n-type and p-type semiconductors – differences and examples. Magnetic properties of substances – paramagnetism, diamagnetism and ferromagnetism, examples.

UNIT-II

Solutions

9 hrs

Types of Solutions – binary – gaseous, liquid and solids, expressing the concentration of a solution of a solid in a liquid – mole fraction, molarity and molality. Solubility, solubility of a gas in a liquid – Henry's law, graph of partial pressure of a gas vs its mole fraction in solution, effect of pressure, temperature, applications of Henry's law.

Solution of liquid in liquid – Raoult's law- statement, mathematical expression, numerical problems, ideal solution – characteristics, graph, non-ideal solution –types – their characteristics and differences, examples, azeotropes – types, examples.

Solution of a solid in a liquid – Raoult's law – colligative properties – relative lowering of vapour pressure, elevation in boiling point, depression in freezing point, graphs for elevation in boiling point and depression in freezing point, SI units for K_b , K_f , osmosis – osmotic pressure, isotonic, hypertonic, hypotonic solutions, reverse osmosis – application in desalination of water. Numerical problems on determination of molar mass using colligative properties.

Abnormal molar mass, van't Hoff factor i , value of i for non-electrolytes and solutes that associate or dissociate in dilute solution.

Guidelines For Setting Ll Puc C

Padhraic Smyth



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