

# Arrhenius' concept based on water

Arrhenius, 1880s:

**Acids form hydrogen ions  $H^+(H_2O)_n$  in aqueous solution.**

**Bases form hydroxide ions in aqueous solution.**

Examples of Arrhenius acids (in water):  $HCl$ ,  $H_2SO_4$ , etc.

Examples of Arrhenius bases (in water):  $NaOH$ ,  $NH_3$ , etc.

Arrhenius definitions only apply to aqueous solutions.

A general Arrhenius acid-base reaction is the reaction between  $H^+$  and  $OH^-$  to produce water.

## A Neutralization Reaction



# Guided Reading And Study Workbook Chapter 20 Acids And Bases

**David Baud**



## **Guided Reading And Study Workbook Chapter 20 Acids And Bases:**

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