

Stoichiometry Guided Practice

Molar to Molar Stoichiometry:



a. What is the mole ratio of Ca(OH)_2 to water?

$$\frac{1 \text{ mol Ca(OH)}_2}{2 \text{ mol H}_2\text{O}}$$

b. How many moles of Ca(OH)_2 will be produced if 4 moles of water are also produced?

$$4 \text{ mol H}_2\text{O} \times \frac{1 \text{ mol Ca(OH)}_2}{2 \text{ mol H}_2\text{O}} = 2 \text{ mol Ca(OH)}_2$$

c. What is the mole ratio of Ca(OH)_2 to water?

$$\frac{1 \text{ mol Ca(OH)}_2}{2 \text{ mol H}_2\text{O}}$$

d. How many moles of water are produced if 4 moles of Ca(OH)_2 are used in the reaction?

$$4 \text{ Ca(OH)}_2 \times \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol Ca(OH)}_2} = 8 \text{ mol H}_2\text{O}$$



a. What is the mole ratio of water to CaCO_3 ?

$$\frac{3 \text{ mol H}_2\text{O}}{1 \text{ mol CaCO}_3}$$

b. How many moles of H_2O will be produced, if 50000 moles of CaCO_3 are produced?

$$50,000 \text{ mol CaCO}_3 \times \frac{3 \text{ mol H}_2\text{O}}{1 \text{ mol CaCO}_3} = 150,000 \text{ mol H}_2\text{O}$$

c. What is the mole ratio of CaCO_3 to NH_3 ?

$$\frac{1 \text{ mol CaCO}_3}{2 \text{ mol NH}_3}$$

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ACT



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