

Chemical Reaction Example Guided Practice Problems

Directions: Write balanced chemical equations for each of the following reactions. Be sure to include phases of matter.

1. Solid iron (III) oxide reacts with gaseous hydrogen chloride to form iron (III) chloride and water vapor.
2. Nitric acid dissolved in water reacts with solid sodium carbonate to form liquid water and carbon dioxide gas and sodium nitrate dissolved in water.

Convert these equations to sentences

3. $\text{Fe}_2\text{O}_3 + \text{O}_{2(g)} \rightarrow \text{Fe}_2\text{O}_{3(g)}$
4. $\text{Cu}_{2(s)} + \text{AgNO}_{3(aq)} \rightarrow \text{Ag}_{2(s)} + \text{Cu}(\text{NO}_3)_{2(aq)}$
5. $\text{NO}_{2(g)} \rightarrow \text{N}_{2(g)} + \text{O}_{2(g)}$

Additional Practice

6. Silver nitrate solution and potassium chromate solution, when mixed together, produce solid silver chromate and potassium nitrate solution.
7. Phosphoric acid solution reacts with barium chloride solution to produce solid barium phosphate and hydrochloric acid solution.
8. Solid magnesium burns in air to produce magnesium oxide, a solid white powder.
9. When solid sulfur is exposed to gaseous fluorine, sulfur hexafluoride gas is formed.
10. When exposed to manganese (IV) oxide, hydrogen peroxide rapidly breaks down into water and oxygen gas.
11. Solid sodium carbonate reacts with hydrochloric acid solution to produce carbon dioxide gas, water and sodium chloride solution.

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Wolfgang Guggemos



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