Chapter 6: Chemical Bonding Guided Notes

B. Why do Atoms Combine?

- A. Chemical Formula
 - The chemical formula shows:
 - a. The _______ in the compound.
 - b. The ratio of their

← H₂O---



- B. Chemical Bond
 - Strong attractive force between ______ or ions in a molecule or compound.
 - 2. Formed by:
 - electrons, (e-) by losing or gaining electrons.
 - electrons (e-)
- C. Stability
 - 1. Octet Rule
 - a. Most atoms form bonds in order to have _______ valence electrons (e-).
 - b. outer energy level.
 - c. Want to be like the Noble Gases (group 18).
 - Stability is the driving force behind bond formation. All atoms bond to become stable.
 - 3. Transferring e-



4. Sharing e-



at. Kinds of Chemical Bonds

- A. Jonic Bond
 - Attraction between 2 charged ions.



- 2. Ions are ______atoms.
- A ________ is a ______ charged ion. Forms when an atom loses an electron. Formed by
- 4. A ______ is a _____ charged ion. Forms when an atom gains an electron. Formed by ______.
- Formed by electrons (e-) from a to a
- lons form a 3-D crystal lattice.



Guided Notes Ionic Bonds Chapter 6

Rachel S Tattersall

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Guided Notes Ionic Bonds Chapter 6 Book Review: Unveiling the Power of Words

In a world driven by information and connectivity, the ability of words has be evident than ever. They have the capacity to inspire, provoke, and ignite change. Such is the essence of the book **Guided Notes Ionic Bonds Chapter 6**, a literary masterpiece that delves deep into the significance of words and their affect our lives. Written by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we will explore the book is key themes, examine its writing style, and analyze its overall impact on readers.

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