



CHEMISTRY

CHEAT SHEET

EQUILIBRIUM

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Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the symbol \rightleftharpoons .

<p>What is Equilibrium?</p> <p>Equilibrium is a state of balance between two or more opposing processes.</p> <p>In a chemical reaction, it is the state where the concentrations of the reactants and products remain constant over time.</p> <p>Equilibrium is a dynamic process, meaning that the reaction is still occurring, but the rates of the forward and reverse reactions are equal.</p>	<p>Equilibrium Constant</p> <p>The equilibrium constant (K) is a measure of the extent to which a reaction proceeds. It is defined as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	
<p>Equilibrium Constant, K_c</p> <p>K_c is the equilibrium constant based on concentrations. It is used for reactions in solution or in the gas phase.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p>Equilibrium Constant, K_p</p> <p>K_p is the equilibrium constant based on partial pressures. It is used for reactions in the gas phase.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K_p = \frac{(P_C)^c (P_D)^d}{(P_A)^a (P_B)^b}$	
<p>Equilibrium Constant, K</p> <p>K is the equilibrium constant based on activities. It is used for reactions in any phase.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K = \frac{a_C^c a_D^d}{a_A^a a_B^b}$	<p>Relationship between K_c and K_p</p> <p>The relationship between K_c and K_p is given by the equation:</p> $K_p = K_c (RT)^{\Delta n}$ <p>where R is the gas constant, T is the temperature in Kelvin, and Δn is the change in the number of moles of gas.</p>	
<p>Le Chatelier's Principle: the effect of changing conditions on equilibrium</p>		
<p>Concentration</p>	<p>Increasing the concentration of a reactant shifts the equilibrium to the right.</p> <p>Increasing the concentration of a product shifts the equilibrium to the left.</p>	<p>$N_2 + 3H_2 \rightleftharpoons 2NH_3$</p> <p>Increasing the concentration of N_2 shifts the equilibrium to the right, producing more NH_3.</p>
<p>Pressure</p>	<p>Increasing the pressure shifts the equilibrium to the side with fewer moles of gas.</p> <p>Decreasing the pressure shifts the equilibrium to the side with more moles of gas.</p>	<p>$N_2 + 3H_2 \rightleftharpoons 2NH_3$</p> <p>Increasing the pressure shifts the equilibrium to the right, producing more NH_3.</p>
<p>Temperature</p>	<p>Increasing the temperature shifts the equilibrium to the endothermic side.</p> <p>Decreasing the temperature shifts the equilibrium to the exothermic side.</p>	<p>$N_2 + 3H_2 \rightleftharpoons 2NH_3$</p> <p>Increasing the temperature shifts the equilibrium to the left, producing less NH_3.</p>

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