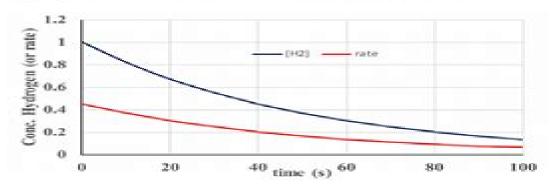
Chemical Kinetics

 The rate of chemical reaction – the change in the amounts of reagents (or products) divided by the change in time.

For the reaction of hydrogen and iodine, $H_2(g) + I_3(g) \rightarrow 2 HI(g)$, the reaction rate is defined as

$$_{\text{ater}} - \frac{\Delta[H_2]}{\Delta t} = - \frac{\Delta[I_2]}{\Delta t} = + \frac{1}{2} \frac{\Delta[HI]}{\Delta t}$$

Rate= Δt Δt Δt Note that reaction rate is a positive number, therefore, negative sign must be added when reagents are used (because their concentrations decrease) and positive sign when products are used (because concentrations of products increase). Factor $\frac{1}{2}$ is due to stoichiometry because two moles of HI are made from one mole of H₂ and I₂. Concentration of hydrogen and the reaction rate (magnified 25 times) for the above reaction, is shown below



 The instantaneous reaction rate - this is the slope of the concentration curve plotted against the reaction time. For general reaction: aA + bB → cC + dD, the reaction rate is defined as:

$$Rate = -\frac{1}{a} \frac{\Delta[A]}{\Delta I} = -\frac{1}{b} \frac{\Delta[B]}{\Delta I} = +\frac{1}{c} \frac{\Delta[C]}{\Delta I} = +\frac{1}{d} \frac{\Delta[D]}{\Delta I}$$

- Measuring reaction rates → determine concentration of reagents or products as a function of time, using suitable analytical techniques, for example, absorption of light or chromatography.
- The rate laws: The reaction rate depends on concentration of reagents. Let's consider reaction A → products (for example, decomposition of sucrose into glucose and

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