

Equilibrium Constants in Terms of Pressure, K_p

When the reactants and products are gases, the equilibrium constant can be expressed in terms of partial pressures (K_p).



The relationship between K_p and K_c is given by the following equation:

$$K_p = K_c (RT)^{\Delta n}$$

$$\Delta n = \sum (\text{product coefficients}) - \sum (\text{reactant coefficients})$$

Example: $3\text{H}_2(g) + \text{N}_2(g) \rightleftharpoons 2\text{NH}_3(g)$ $K = 5.80 \times 10^2$, $K_p = ?$

$$\Delta n = 2 - (3 + 1) = -2 \quad K_p = 5.80 \times (0.08206 \times 273)^{-2} = 4.34 \times 10^3$$

- If the number of gas molecules does not change in the reaction, $K_p = K_c$ because $\Delta n = 0$.

K for Heterogeneous Equilibria

The concentrations of solids, liquids, and the solvent are omitted from equilibrium expressions.



Equilibrium Constants and Units

The more accurate expression for the equilibrium constant is given by using the activities of reactants and products.

$$K = \frac{(a_c)^c (a_d)^d}{(a_a)^a (a_b)^b}$$

For most undergraduate general chemistry courses, it is assumed that $a \approx 1$ since the differences are small and can be neglected. Because activities have no units, K is traditionally shown without units.

The Direction of Reaction and Quotient (Q)

To predict the direction of reaction before an equilibrium is reached, reaction quotient is used. It is similar to the expression for K , however, the **actual concentrations are used**.



- $Q < K$ Reaction has a tendency to form more products.
- $Q > K$ Reaction has a tendency to form more reactants.
- $Q = K$ Reaction is at equilibrium.

$$Q_c = \frac{[C]^c [D]^d}{[A]^a [B]^b} \quad Q_p = \frac{P_C^c P_D^d}{P_A^a P_B^b}$$

- If any reactant or product is missing in the initial mixture, the reaction will shift in the direction forming some of that component.

Le Chatelier's Principle

If a chemical system at equilibrium is disturbed, the system shifts in a direction that minimizes the disturbance. This can be done in different ways.

Changing Concentrations at Equilibrium

- Adding a reactant produces more product by consuming some of the reactant.
- Adding a product produces more reactant by consuming some of the product.



Changing The Reaction Volume

- Decreasing the volume shifts the reaction to in the direction with less gas molecules.
- Increasing the volume shifts the reaction to in the direction with more gas molecules.



Changing The Pressure

- The pressure can be changed in a few ways: 1) By decreasing the volume (see above), 2) By adding more volume to the mixture (see above), 3) By adding an inert gas.
- Adding an inert gas does not cause a shift in equilibrium because it does not alter the partial pressures of any of the reacting components.

Changing The Temperature

Know the sign of ΔH is different than what we write for heat!

For an endothermic reaction, treat the heat as a product.



For an exothermic reaction, treat the heat as a reactant.



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Table of Contents General Chemistry Ii Cheat Guide

1. Understanding the eBook General Chemistry Ii Cheat Guide
 - The Rise of Digital Reading General Chemistry Ii Cheat Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying General Chemistry Ii Cheat Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an General Chemistry Ii Cheat Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from General Chemistry Ii Cheat Guide
 - Personalized Recommendations
 - General Chemistry Ii Cheat Guide User Reviews and Ratings
 - General Chemistry Ii Cheat Guide and Bestseller Lists
5. Accessing General Chemistry Ii Cheat Guide Free and Paid eBooks
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- General Chemistry Ii Cheat Guide eBook Subscription Services
- General Chemistry Ii Cheat Guide Budget-Friendly Options
- 6. Navigating General Chemistry Ii Cheat Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - General Chemistry Ii Cheat Guide Compatibility with Devices
 - General Chemistry Ii Cheat Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of General Chemistry Ii Cheat Guide
 - Highlighting and Note-Taking General Chemistry Ii Cheat Guide
 - Interactive Elements General Chemistry Ii Cheat Guide
- 8. Staying Engaged with General Chemistry Ii Cheat Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers General Chemistry Ii Cheat Guide
- 9. Balancing eBooks and Physical Books General Chemistry Ii Cheat Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection General Chemistry Ii Cheat Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine General Chemistry Ii Cheat Guide
 - Setting Reading Goals General Chemistry Ii Cheat Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of General Chemistry Ii Cheat Guide
 - Fact-Checking eBook Content of General Chemistry Ii Cheat Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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