

Which of these are **true** and which are **false** about acid-base reactions?

- In general, an acid-base reaction will be favored when a weaker acid and a weaker base combine to form a stronger acid and a stronger base  
*Other way around: will be favored when a stronger acid and a stronger base combine to give a weaker acid and a weaker base.*  
T F
- Generally, strong acids have weak conjugate bases  
*True - and weak acids have strong conjugate bases*  
T F
- Strong acids tend to have high  $pK_a$  values  
*Strong acids have low (strongly negative)  $pK_a$  values - e.g.  $HCl$  has a  $pK_a$  of  $-8$*   
T F
- The greater the acidity of a molecule, the weaker a base it is  
*Not quite correct. The stronger the acid, the weaker is its conjugate base.*  
T F
- The inductive effect of electronegative atoms near a carboxylic acid group increases the acidity of the compound.  
T F
- Any factor which stabilizes negative charge generally tends to decrease the acidity of an acid  
*False - increases acidity, not decreases.*  
T F
- The  $pK_a$  of  $NH_4$  is 10 and the  $pK_a$  of  $H_2O$  is 14, which means that  $H_2O$  is a stronger acid than  $H_2O$   
*False -  $NH_4$  is a stronger acid by a factor of 4  $pK_a$  units (about  $10^4$ )*  
T F
- The acidity of a molecule can be increased by resonance stabilization of its conjugate base.  
T F
- $sp$ -hybridized C-H bonds are more acidic than  $sp^3$ -hybridized C-H bonds since the resulting negative charge is closer to the nucleus  
T F
- Acid-base reactions where a stronger base is formed from a weaker base never occur.  
*They do occur and they can be important. E.g.  $NH_3$  can deprotonate  $H_2O$  even though it is a weaker base by a factor of  $10^4$  compared to  $HO^-$ . Even a difference in 9-10  $pK_a$  units can be significant for moving some reactions forward.*  
T F

# General Chemistry Acid Base Practice Problems

**Kenneth W. Raymond**



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## **Table of Contents General Chemistry Acid Base Practice Problems**

1. Understanding the eBook General Chemistry Acid Base Practice Problems
  - The Rise of Digital Reading General Chemistry Acid Base Practice Problems
  - Advantages of eBooks Over Traditional Books
2. Identifying General Chemistry Acid Base Practice Problems
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an General Chemistry Acid Base Practice Problems
  - User-Friendly Interface
4. Exploring eBook Recommendations from General Chemistry Acid Base Practice Problems

- Personalized Recommendations
- General Chemistry Acid Base Practice Problems User Reviews and Ratings
- General Chemistry Acid Base Practice Problems and Bestseller Lists
- 5. Accessing General Chemistry Acid Base Practice Problems Free and Paid eBooks
  - General Chemistry Acid Base Practice Problems Public Domain eBooks
  - General Chemistry Acid Base Practice Problems eBook Subscription Services
  - General Chemistry Acid Base Practice Problems Budget-Friendly Options
- 6. Navigating General Chemistry Acid Base Practice Problems eBook Formats
  - ePub, PDF, MOBI, and More
  - General Chemistry Acid Base Practice Problems Compatibility with Devices
  - General Chemistry Acid Base Practice Problems Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of General Chemistry Acid Base Practice Problems
  - Highlighting and Note-Taking General Chemistry Acid Base Practice Problems
  - Interactive Elements General Chemistry Acid Base Practice Problems
- 8. Staying Engaged with General Chemistry Acid Base Practice Problems
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers General Chemistry Acid Base Practice Problems
- 9. Balancing eBooks and Physical Books General Chemistry Acid Base Practice Problems
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection General Chemistry Acid Base Practice Problems
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine General Chemistry Acid Base Practice Problems
  - Setting Reading Goals General Chemistry Acid Base Practice Problems
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of General Chemistry Acid Base Practice Problems

- Fact-Checking eBook Content of General Chemistry Acid Base Practice Problems
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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