

**Review for the Full Year General Chemistry
American Chemical Society Exam**

(a) Thermodynamics, (b) Transition Metals, (c) Redox and Electrochemistry, (d) Nuclear and (e) Organic

THERMODYNAMICS

84. A 10.0 g sample of silver is heated to 100.0 °C and then added to 20.0 g of water at 23.0 °C in an insulated calorimeter. At thermal equilibrium the temperature of the system was measured as 25.0 °C. What is the specific heat of silver? [specific heat = $4.2 \text{ J g}^{-1} \text{ }^{\circ}\text{C}^{-1}$]

- (A) 0.11 J/g K (C) 17 J/g K
(B) 0.22 J/g K (D) 34 J/g K

85. What is ΔH in kJ for the reaction:



- (A) -1830 kJ (B) -2284 kJ
(C) -2510 kJ (D) -1605 kJ

86. Given these equations:



calculate the heat of formation of $\text{SO}_3(\text{g})$.

- (A) -500 kJ mol⁻¹ (C) +100 kJ mol⁻¹
(B) -400 kJ mol⁻¹ (D) +200 kJ mol⁻¹

87. Given:



How much energy is produced in the combustion of 28 g of CO? [Atomic Masses: C 12.0 g mol⁻¹; O 16.0 g mol⁻¹]

- (A) 88 kJ (B) 109 kJ
(C) 175 kJ (D) 284 kJ

88. Calculate the value of ΔH (in kJ mol⁻¹) for the reaction:



Bond Energies (kJ mol ⁻¹)	
H-H	435
N=N	945 (in N ₂)
N-H	389

- (A) 2340 kJ of heat absorbed
(B) 213 kJ of heat absorbed
(C) 2340 kJ of heat evolved
(D) 83 kJ of heat evolved

89. Which change is likely to be accompanied by the greatest increase in entropy?



90. For which process is the entropy change per mole the largest at constant temperature?



91. In which process is entropy decreased?

- (A) dissolving sugar in water
(B) expanding a gas
(C) evaporating a liquid
(D) freezing water

92. When $\text{Al}_2\text{O}_3(\text{s})$ is formed from the elements at standard conditions, the values of ΔH° and ΔG° at 298 K are -1676 kJ mol⁻¹ and -1577 kJ mol⁻¹, respectively. The standard entropy of formation per mole, in joules per degree, will be

- (A) -332 (B) -157 (C) -93.3
(D) -0.0933 (E) +15.7

93. Vaporization of a liquid is an example of a process for which

- (A) ΔH , ΔS , and ΔG are positive at all temperatures.
(B) ΔH and ΔS are positive.
(C) ΔG is negative at low temperatures, positive at high temperatures.
(D) $\Delta H = \Delta S$

94. A particular chemical reaction has a negative ΔH and negative ΔS . Which statement is correct?

- (A) The reaction is spontaneous at all temperatures.
(B) The reaction is nonspontaneous at all temperatures.
(C) The reaction becomes spontaneous as temperature increases.
(D) The reaction becomes spontaneous as temperature decreases.

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1. Understanding the eBook General Chemistry 1 Standardized Final Exam Review
 - The Rise of Digital Reading General Chemistry 1 Standardized Final Exam Review
 - Advantages of eBooks Over Traditional Books
2. Identifying General Chemistry 1 Standardized Final Exam Review
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an General Chemistry 1 Standardized Final Exam Review
 - User-Friendly Interface
4. Exploring eBook Recommendations from General Chemistry 1 Standardized Final Exam Review
 - Personalized Recommendations
 - General Chemistry 1 Standardized Final Exam Review User Reviews and Ratings
 - General Chemistry 1 Standardized Final Exam Review and Bestseller Lists

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 - General Chemistry 1 Standardized Final Exam Review Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
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 - General Chemistry 1 Standardized Final Exam Review Enhanced eBook Features
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 - Highlighting and Note-Taking General Chemistry 1 Standardized Final Exam Review
 - Interactive Elements General Chemistry 1 Standardized Final Exam Review
8. Staying Engaged with General Chemistry 1 Standardized Final Exam Review
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers General Chemistry 1 Standardized Final Exam Review
9. Balancing eBooks and Physical Books General Chemistry 1 Standardized Final Exam Review
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection General Chemistry 1 Standardized Final Exam Review
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine General Chemistry 1 Standardized Final Exam Review
 - Setting Reading Goals General Chemistry 1 Standardized Final Exam Review
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of General Chemistry 1 Standardized Final Exam Review
 - Fact-Checking eBook Content of General Chemistry 1 Standardized Final Exam Review
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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