

Chem 101 Practice Test: Thermodynamics 3.5 – 3.6, 11.1**Equations**

$$\Delta G = \Delta H - T\Delta S$$

$$\Delta G = \Delta H_{\text{sys}} + P\Delta V_{\text{sys}}$$

$$\Delta G = \Delta H_{\text{sys}}$$

$$\Delta H_{\text{sys}} = \sum \text{products} - \sum \text{reactants}$$

$$\Delta S = \Delta S_{\text{sys}} + \Delta S_{\text{surr}}$$

Properties of water

$$\Delta_{\text{vap}} H^\circ = 40.7 \text{ kJ/mol}$$

$$\Delta_{\text{fus}} H^\circ = 6.01 \text{ kJ/mol}$$

$$\Delta_{\text{sub}} H^\circ = 47.05 \text{ kJ/mol}$$

$$\Delta_{\text{vap}} S^\circ = 109 \text{ J/mol}\cdot\text{K}$$

$$\Delta_{\text{fus}} S^\circ = 22 \text{ J/mol}\cdot\text{K}$$

Multiple choice

Multiple choice and select-all-that-apply questions follow. One or more may be correct.

(1) spontaneous

(2) exothermic

(3) faster

(4) exothermic

(5) faster exothermic

(6) specific heat

(7) faster heating

(8) heat of vaporization

(9) heat of fusion

- ☒ 1. spontaneity of thermodynamic favorability depends on the temperature of the system.
- ☒ 2. all work is reversible
- ☒ 3. spontaneity of thermodynamic favorability depends on the temperature of the system.
- ☒ 4. thermodynamic favorability of thermodynamic favorability depends on the temperature of the system.
- ☒ 5. heat of fusion is a constant pressure.
- ☒ 6. the enthalpy change is constant for thermodynamic favorability.
- ☒ 7. the energy of fusion is constant for thermodynamic favorability.

Multiple choice (select all that apply)

8. Which of the following are spontaneous processes?
- ☒ a. A system that is at equilibrium with its surroundings.
- ☒ b. A system that is at equilibrium with its surroundings.
- ☐ c. A system that is at equilibrium with its surroundings.
- ☐ d. A system that is at equilibrium with its surroundings.
9. Which of the following are spontaneous processes?
- ☒ a. exothermic process
- ☒ b. endothermic process
- ☒ c. exothermic process
- ☒ d. endothermic process
10. In which of the following is the system at equilibrium with its surroundings?
- ☒ a. exothermic process
- ☒ b. endothermic process
- ☒ c. exothermic process
- ☒ d. endothermic process

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